

Enthalpy change for the thermal decomposition of potassium hydrogen carbonate =
 $5.38 \text{ KJ/mol} + 0.32 \text{ KJ/mol} = 5.7 \text{ KJ/mol}$

Enthalpy change for the thermal decomposition of potassium hydrogen carbonate = 5.7 KJ/mol

4. When KHCO_3 is heated, it decomposes according to the following chemical reaction:



Apply Hess' Law and use the results from questions 2b and 3b to calculate the enthalpy change for the thermal decomposition of potassium hydrogen carbonate.

Standard Enthalpy Change Of Potassium Hydrogen Carbonate

David Kirk



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