Chemical Equilibrium

Multiple Choice Practice

- Chemical reactions are in a state of dynamic equilibrium only when
 - A. the rate of the forward reaction equals that of the backward reaction
 - B. the concentration of reactants and products are equal
 - C. the activation energies of the forward and backward reactions are equal
 - D. the reaction involves no enthalpy change.
- Under the conditions used industrially, ethene and steam react as follows:

$$C_2H_4(g) + H_2O(g) \rightarrow C_2H_2OH(g)$$

 $\Delta H = -46 \text{ kJ mol}^{-1}$

Which set of conditions would give the best yield of ethanol at equilibrium?

- A. high temperature, low pressure
- high temperature, high pressure
- C. low temperature, high pressure
- D. low temperature, low pressure.
- N₂ (g) + 3H₂ (g) → 2NH₃ (g)

In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because

- A. the ammonia is constantly removed
- B. an iron catalyst is used
- C. high pressure is used.
- D. the temperature is held at above 500°C.

 The equation refers to the preparation of methanol from synthesis gas.

CO(g) +
$$2H_2(g) \leftrightarrow CH_3OH(g)$$

 $\Delta H = -91 \text{ kJ mol}^{-1}$

The formation of methanol is favoured by

- high pressure and low temperature
- B. high pressure and high temperature
- C. low pressure and low temperature
- D. low pressure and high temperature.
- ICI(I) + Cl₂(g) ↔ ICI₃(s)
 AH = -46 kJ mol⁻¹

Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?

| | Temperature | Pressure |
|---|-------------|----------|
| A | decrease | decrease |
| В | decrease | increase |
| C | increase | decrease |
| D | increase | increase |

- When a reversible chemical reaction is at equilibrium.
 - A. the concentration of reactants and products remain equal
 - B. the forward reaction is unable to continue
 - C. the concentration of reactants and products remain constant
 - D. the forward and reverse reactions proceed at different rates.

Static Equilibrium Multiple Choice

Bryan McCabe, Marina Pantazidou, Declan Phillips Static Equilibrium Multiple Choice:

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