

**Heat of Formation Worksheet**

Use a standard enthalpies of formation table to determine the change in enthalpy for each of these reactions.



$$[(-411.0) + (-241.8)] - [(-426.7) + (-92.3)] = -133.8 \text{ kJ}$$



$$[2(-393.5)] - [2(-110.5)] = -566 \text{ kJ}$$



$$[(-393.5) + 2(-285.8)] - [(-74)] = -891.1 \text{ kJ}$$



$$[2(-285.8) + 2(-296.1)] - [2(-20.1)] = -1123.6 \text{ kJ}$$



$$[2(+33.9)] - [2(+90.4)] = -113 \text{ kJ}$$

Compound	$\Delta H_f$ (kJ/mol)	Compound	$\Delta H_f$ (kJ/mol)
$\text{CH}_4\text{(g)}$	-74.8	$\text{HCl(g)}$	-92.3
$\text{CO}_2\text{(g)}$	-393.5	$\text{H}_2\text{O(g)}$	-241.8
$\text{NaCl(s)}$	-411.0	$\text{SO}_2\text{(g)}$	-296.1
$\text{H}_2\text{O(l)}$	-285.8	$\text{NH}_4\text{Cl(s)}$	-315.4
$\text{H}_2\text{S(g)}$	-20.1	$\text{NO(g)}$	+90.4
$\text{H}_2\text{SO}_4\text{(l)}$	-811.3	$\text{NO}_2\text{(g)}$	+33.9
$\text{MgSO}_4\text{(s)}$	-1278.2	$\text{SnCl}_4\text{(l)}$	-545.2
$\text{MnO(s)}$	-384.9	$\text{SnO(s)}$	-286.2
$\text{MnO}_2\text{(s)}$	-519.7	$\text{SnO}_2\text{(s)}$	-580.7
$\text{NaCl(s)}$	-411.0	$\text{SO}_3\text{(g)}$	-296.1
$\text{NaF(s)}$	-569.0	$\text{SO}_3\text{(l)}$	-395.2
$\text{NaOH(s)}$	-426.7	$\text{ZnO(s)}$	-348.0
$\text{NH}_3\text{(g)}$	-46.2	$\text{ZnS(s)}$	-202.9

# Standard Heat Of Formation Answer Key

**Disha Experts**



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### **Table of Contents Standard Heat Of Formation Answer Key**

1. Understanding the eBook Standard Heat Of Formation Answer Key
  - The Rise of Digital Reading Standard Heat Of Formation Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Standard Heat Of Formation Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in a Standard Heat Of Formation Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Standard Heat Of Formation Answer Key
  - Personalized Recommendations
  - Standard Heat Of Formation Answer Key User Reviews and Ratings
  - Standard Heat Of Formation Answer Key and Bestseller Lists

5. Accessing Standard Heat Of Formation Answer Key Free and Paid eBooks
  - Standard Heat Of Formation Answer Key Public Domain eBooks
  - Standard Heat Of Formation Answer Key eBook Subscription Services
  - Standard Heat Of Formation Answer Key Budget-Friendly Options
6. Navigating Standard Heat Of Formation Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Standard Heat Of Formation Answer Key Compatibility with Devices
  - Standard Heat Of Formation Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Standard Heat Of Formation Answer Key
  - Highlighting and Note-Taking Standard Heat Of Formation Answer Key
  - Interactive Elements Standard Heat Of Formation Answer Key
8. Staying Engaged with Standard Heat Of Formation Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Standard Heat Of Formation Answer Key
9. Balancing eBooks and Physical Books Standard Heat Of Formation Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Standard Heat Of Formation Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Standard Heat Of Formation Answer Key
  - Setting Reading Goals Standard Heat Of Formation Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Standard Heat Of Formation Answer Key
  - Fact-Checking eBook Content of Standard Heat Of Formation Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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