

## Chemical Equilibrium

### Multiple Choice Practice

1. Chemical reactions are in a state of dynamic equilibrium only when
- A. the rate of the forward reaction equals that of the backward reaction.
  - B. the concentration of reactants and products are equal
  - C. the activation energies of the forward and backward reactions are equal
  - D. the reaction involves no enthalpy change.
2. Under the conditions used industrially, ethene and steam react as follows:



Which set of conditions would give the best yield of ethanol at equilibrium?

- A. high temperature, low pressure
  - B. high temperature, high pressure
  - C. low temperature, high pressure
  - D. low temperature, low pressure.
3.  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
- In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because
- A. the ammonia is constantly removed
  - B. an iron catalyst is used
  - C. high pressure is used
  - D. the temperature is held at above  $500^\circ\text{C}$ .

4. The equation refers to the preparation of methanol from synthesis gas.
- $$\text{CO}(\text{g}) + 2\text{H}_2(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g})$$
- $$\Delta H = -91 \text{ kJ mol}^{-1}$$

The formation of methanol is favoured by

- A. high pressure and low temperature
  - B. high pressure and high temperature
  - C. low pressure and low temperature
  - D. low pressure and high temperature.
5.  $\text{ICl}(\text{l}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{ICl}_3(\text{s})$
- $$\Delta H = -46 \text{ kJ mol}^{-1}$$

Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?

	Temperature	Pressure
A	decrease	decrease
B	decrease	increase
C	increase	decrease
D	increase	increase

6. When a reversible chemical reaction is at equilibrium,
- A. the concentration of reactants and products remain equal
  - B. the forward reaction is unable to continue
  - C. the concentration of reactants and products remain constant
  - D. the forward and reverse reactions proceed at different rates.

# Static Equilibrium Multiple Choice Questions

**Arshad Iqbal**



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**Introductory Physics for the Life Sciences: Mechanics (Volume One)** David V. Guerra, 2023-05-09 This textbook provides an accessible introduction to physics for undergraduate students in the life sciences including those majoring in all branches of biology biochemistry and psychology and students working on pre professional programs such as pre medical pre dental and physical therapy The text is geared for the algebra based physics course often named College Physics in the United States The order of topics studied are such that most of the problems in the text can be solved with the methods of Statics or Dynamics That is they require a free body diagram the application of Newton s Laws and any necessary kinematics Constructing the text with a standardized problem solving methodology simplifies this aspect of the course and

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## **Table of Contents Static Equilibrium Multiple Choice Questions**

1. Understanding the eBook Static Equilibrium Multiple Choice Questions
  - The Rise of Digital Reading Static Equilibrium Multiple Choice Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Static Equilibrium Multiple Choice Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Static Equilibrium Multiple Choice Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Static Equilibrium Multiple Choice Questions
  - Personalized Recommendations
  - Static Equilibrium Multiple Choice Questions User Reviews and Ratings
  - Static Equilibrium Multiple Choice Questions and Bestseller Lists
5. Accessing Static Equilibrium Multiple Choice Questions Free and Paid eBooks
  - Static Equilibrium Multiple Choice Questions Public Domain eBooks

- Static Equilibrium Multiple Choice Questions eBook Subscription Services
- Static Equilibrium Multiple Choice Questions Budget-Friendly Options
- 6. Navigating Static Equilibrium Multiple Choice Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Static Equilibrium Multiple Choice Questions Compatibility with Devices
  - Static Equilibrium Multiple Choice Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Static Equilibrium Multiple Choice Questions
  - Highlighting and Note-Taking Static Equilibrium Multiple Choice Questions
  - Interactive Elements Static Equilibrium Multiple Choice Questions
- 8. Staying Engaged with Static Equilibrium Multiple Choice Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Static Equilibrium Multiple Choice Questions
- 9. Balancing eBooks and Physical Books Static Equilibrium Multiple Choice Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Static Equilibrium Multiple Choice Questions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Static Equilibrium Multiple Choice Questions
  - Setting Reading Goals Static Equilibrium Multiple Choice Questions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Static Equilibrium Multiple Choice Questions
  - Fact-Checking eBook Content of Static Equilibrium Multiple Choice Questions
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

## 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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