

Assignment #3 – Compound Names and Formulas
Single and Multi-valent ions

A. Write the correct formula for the following compounds.

- | | |
|--|---|
| 1. gallium oxide <u>Ga_2O_3</u> | 11. silicon oxide <u>SiO_2</u> |
| 2. mercury (I) chloride <u>$HgCl$</u> | 12. aluminum bromide <u>$AlBr_3$</u> |
| 3. tin (IV) bromide <u>$SnBr_4$</u> | 13. tin (II) sulphide <u>SnS</u> |
| 4. indium oxide <u>In_2O_3</u> | 14. germanium bromide <u>$GeBr_4$</u> |
| 5. cadmium sulphide <u>CdS</u> | 15. nickel (II) oxide <u>NiO</u> |
| 6. magnesium oxide <u>MgO</u> | 16. strontium astatide <u>$SrAt_2$</u> |
| 7. potassium nitride <u>K_3N</u> | 17. aluminum oxide <u>Al_2O_3</u> |
| 8. iron (III) sulphide <u>Fe_2S_3</u> | 18. copper (II) oxide <u>CuO</u> |
| 9. gold (III) bromide <u>$AuBr_3$</u> | 19. lawrencium fluoride <u>LrF_3</u> |
| 10. zinc oxide <u>ZnO</u> | 20. actinium oxide <u>Ac_2O_3</u> |

B. Name the following compounds correctly. Only include the Roman Numeral in the name if the first element has 2 or more combining capacities.

- | | |
|--|---|
| 1. $CaBr_2$ <u>Calcium Bromide</u> | 9. $AsCl_5$ <u>Arsonic (V) Chloride</u> |
| 2. K_2O <u>Potassium Oxide</u> | 10. Eu_2O_3 <u>Europium (III) Oxide</u> |
| 3. Fe_2O_3 <u>Iron (III) Oxide</u> | 11. $NdBr_3$ <u>Neodymium Bromide</u> |
| 4. ZnF_2 <u>Zinc Fluoride</u> | 12. Cu_3N_2 <u>Copper (II) Nitride</u> |
| 5. $CuCl$ <u>Copper (I) Chloride</u> | 13. CoS <u>Cobalt (II) Sulphide</u> |
| 6. Ni_2O_3 <u>Nickel (III) Oxide</u> | 14. ZrO_2 <u>Zirconium Oxide</u> |
| 7. $AuCl$ <u>Gold (I) Chloride</u> | 15. Li_2O <u>Lithium Oxide</u> |
| 8. $HgBr$ <u>Mercury (I) Bromide</u> | 16. NaH <u>Sodium Hydride</u> |

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Barry McBride

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