

General Science 3200

Worksheet 4: Writing Formulas for Ionic Compounds

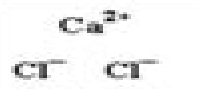
Name: _____

Use the periodic table to find the two ions (positive ion and negative ion) for each compound. Then write the formula.

Remember, the overall charge of the compound is neutral (zero charge). Thus, the overall positive charge must be equal to the overall negative charge.

Example: calcium chloride

Ca^{2+} Cl^{-} (These two charges are not the same, but 2 chloride ions will give an overall negative charge of -2. The positive ion already has a +2 charge.)



Count up the total charge.
+2, -1, -1
This adds up to zero!!

Therefore, $CaCl_2$ is the answer.

Name	Positive Ion	Negative Ion	Formula
1) sodium chloride ex.	Na^{+}	Cl^{-}	$NaCl$
2) potassium bromide			
3) calcium oxide			
4) magnesium sulfide			
5) zinc chloride			
6) sodium sulfide			
7) strontium fluoride			
8) lithium oxide			
9) potassium nitride			
10) sodium phosphide			
11) aluminum chloride			
12) calcium nitride			
13) aluminum oxide			
14) magnesium phosphide			
15) aluminum sulfide			

Skills Worksheet Math Skills Writing Ionic Formulas Answers

McGraw-Hill Staff



Skills Worksheet Math Skills Writing Ionic Formulas Answers:

Science Spectrum Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2003-03 *Prentice Hall Physical Science Concepts in Action Program Planner National Chemistry Physics Earth Science*, 2003-11 Prentice Hall Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Carolina Science and Math** Carolina Biological Supply Company, 2003 **Chemistry** McGraw-Hill Staff, 2001-05-01 **Chemistry Math Concepts** David C. McMillin, 2001 Essential Math for Chemistry Students David W. Ball, 1996 Focuses on the skills needed to survive in general chemistry with worked examples showing how these skills translate into successful chemical problem solving An ideal tool for students lacking in confidence or competency in the essential math skills required for general chemistry

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