

- ④ To find the molecular formula, multiply the empirical formula by the ratio of molar mass divided by empirical mass. In this example it is $66.0/33.0$, which equals 2.
- 4 a Calcium ions are $2+$ and chloride ions are $1-$, so the formula is CaCl_2 .
 b Silver ions are $1+$ and nitrate ions are $1-$, so the formula is AgNO_3 .
 c Copper(II) is Cu^{2+} and phosphate is PO_4^{3-} . For a neutral compound you need three copper ions and two phosphate ions, making charges of $+6$ and -6 . Therefore, the formula is $\text{Cu}_3(\text{PO}_4)_2$.
 d Aluminium is $3+$ and oxide is $2-$, so aluminium oxide is Al_2O_3 .
 ⑤ There are two ways of working out formulae of ionic compounds. One way is to make sure that the total positive charge equals the total negative charge (see part c). The other way, if the charges are not the same, is to swap them round. Copper is $2+$ and phosphate $3-$, so there are three copper ions and two phosphate ions in the formula.
- 5 a As there is one manganese ion to each carbonate ion, the charges have the same numerical value. Carbonate ions are $2-$, so the manganese ions in this compound must be $2+$.
 b The charge on two vanadium ions must equal the charge on three sulfate ions. Each sulfate is -2 , so three sulfate ions have a total charge of -6 . This means that two vanadium ions have a total charge of $+6$, so each has a charge of $+3$.
- 6 a molar mass of $\text{Ca}(\text{OH})_2 = 40.1 + 2 \times (16.0 + 1.0) = 74.1 \text{ g mol}^{-1}$
 b molar mass of $\text{Al}_2(\text{SO}_4)_3 = (2 \times 27.0) + 3 \times [32.1 + (4 \times 16.0)] = 342.3 \text{ g mol}^{-1}$
 c molar mass of $\text{FeSO}_4 \cdot 7\text{H}_2\text{O} = 55.8 + 32.1 + (4 \times 16.0) + 7 \times (2.0 + 16.0) = 277.9 \text{ g mol}^{-1}$
 ⑥ In calculating molar mass, make sure that you use the relative atomic masses (the larger number), not the atomic numbers, from the periodic table. Molar mass has the units of g mol^{-1} . Relative molecular mass has no units. The numbers are the same.
- 7 The balanced equations are:
 a $\text{P}_4 + 5\text{O}_2 \rightarrow \text{P}_4\text{O}_{10}$ or $2\text{P}_2\text{O}_5$
 ⑥ Phosphorus(V) oxide exists as the dimer P_4O_{10} , but $2\text{P}_2\text{O}_5$ on the right of the equation would be acceptable.
 b $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 c $2\text{Mg}(\text{NO}_3)_2 \rightarrow 2\text{MgO} + 4\text{NO}_2 + \text{O}_2$
 ⑥ In part c, balance the oxygen atoms last. Remember that the number of oxygen atoms on the right-hand side of the equation must be an even number, because $\text{Mg}(\text{NO}_3)_2$ contains an even number of oxygen atoms.
 d $2\text{LiOH} + \text{H}_2\text{SO}_4 \rightarrow \text{Li}_2\text{SO}_4 + 2\text{H}_2\text{O}$
 e $2\text{Fe}^{2+} + 2\text{I}^- \rightarrow 2\text{Fe}^{3+} + \text{I}_2$
 ⑥ Make sure that this equation balances for charge. Having one Fe^{2+} ion on the left-hand side and one Fe^{3+} on the right-hand side balances the number of Fe ions, but the charge on the left would not equal the charge on the right.
 f $2\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$ or $\text{C}_8\text{H}_{18} + 12\frac{1}{2}\text{O}_2 \rightarrow 8\text{CO}_2 + 9\text{H}_2\text{O}$
 ⑥ The only source of oxygen on the left-hand side of the equation is O_2 , so the number of oxygen atoms on the right-hand side must be an even number. Therefore, there must be an even number of H_2O molecules on the right. This requires two C_8H_{18} molecules on the left.
- 8 For equations a–c, the simplest way is to realise that the component ions of the compound combine to give the precipitate. All other ions are spectators.

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web it stars charlie sheen kiefer sutherland chris o donnell oliver platt tim curry and rebecca de mornay 2 the film is loosely based on the 1844 novel the three musketeers les trois mousquetaires by alexandre dumas it recounts the adventures of d artagnan on his quest to join the three title characters in becoming a musketeer

the three musketeers tv series 2014 imdb - Feb 23 2022

web the three musketeers with jung yong hwa lee jin wook dong kun yang jung hae in park dal hyan wishing to become a soldier becomes involved with the crown prince and his faithful warriors who called themselves the three musketeers in a joseon ruled by a weak king and in conflict with the quing dynasty

the three musketeers d artagnan wikipedia - Mar 27 2022

web plot d artagnan a young man from gascony is almost killed trying to stop a young woman from being kidnapped by a mysterious older woman and her henchmen he then travels to paris to join the musketeers of the guard d artagnan visits capitaine de tréville commander of the musketeers at its headquarters

the three musketeers wikipedia - Oct 14 2023

web chapter 1 les trois présents de m d artagnan père the three musketeers french les trois mousquetaires le tʁwa muskøtɛːʁ is a french historical adventure novel written in 1844 by french author alexandre dumas as with some of his other works he wrote it in collaboration with ghostwriter auguste maquet

the three musketeers rotten tomatoes - Jan 05 2023

web the king s musketeers athos matthew macfadyen porthos ray stevenson and aramis luke evans have been in a rut for a year since being betrayed by milady de winter milla jovovich in venice