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Covalent Bonding Practice Worksheet

Explanation: Covalent compounds are compounds between nonmetal atoms. The atoms "start" with a number of electrons equal to their group number in the periodic table. Then they share electrons so that all the atoms have filled shells with a "magic number" of electrons. The magic number is 2 for hydrogen and 8 for C, N, O, F, Si, P, S, Cl, Se, Br and I. Remember that compounds made with a metal atom and a nonmetal atom make an ionic compound, not a covalent one, and compounds of metals are also different, these have metallic bonds and do not require a strict ratio.

Instructions: Fill in the blanks in the table. The first row is done for you; study the first row carefully before you start, especially the difference between different boxes (remember there are many ways to draw the same compound). Some rows have hints. If a box has an X you don't have to fill in that box.

Helpful Tips: 1) All atoms start with a number of outer shell electrons = their group number. You cannot change the element's group number. 2) If you have to choose between two shapes, the symmetrical, balanced one is usually best. 3) Try to figure the structures out yourself, don't cheat and use the Internet. ☺

Compound Formula	Full electronic structure	Lewis structure	Dot-cross diagram	Structural formula	3-d formula
H ₂ O		$\text{H} : \ddot{\text{O}} : \text{H}$	$\text{H}^{\times} \ddot{\text{O}}^{\times} \text{H}$	H-O-H	
NH ₃		$\begin{array}{c} \cdot\cdot \\ \text{H} : \text{N} : \text{H} \\ \\ \text{H} \end{array}$			
CH ₃ Br			$\begin{array}{c} \text{H} \\ \times \\ \text{H} : \text{C} : \text{H} \\ \times \\ : \text{Br} : \\ \cdot\cdot \end{array}$		

The Nature Of Covalent Bonding 16 1 Worksheet

Team Prabhat



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multiple bonds which can be determined by calculating the energy in the usual way as though the bonds were single but of the experimental length This book discusses as well the reduction of the lone pair bond weakening effect through the formation of multiple bonds The final chapter deals with the relative roles of principles and practice in the teaching of inorganic and general chemistry This book is a valuable resource for chemists and students

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