

8.2**THE NATURE OF COVALENT BONDING****Section Review****Objectives**

- State a rule that usually tells how many electrons are shared to form a covalent bond.
- Describe how electrons that form bonds are used.
- Predict when two atoms are likely to be joined by a double or a triple covalent bond.
- Distinguish between a single covalent bond and other covalent bonds.
- Describe how the strength of a covalent bond is related to its bond dissociation energy.
- Describe how resonance structures explain bonding.
- Identify some exceptions to the octet rule.

Vocabulary

- | | | |
|------------------------|----------------------------|----------------------------|
| • single covalent bond | • double covalent bonds | • polyatomic ion |
| • structural formulas | • triple covalent bonds | • bond dissociation energy |
| • unshared pairs | • coordinate covalent bond | • resonance structures |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are presented in this section. Each blank can be completed with a term, short phrase, or number.

- When atoms share electrons to gain the ____1____ configuration of a noble gas, the bonds formed are ____2____. A ____3____ pair of valence electrons constitutes a ____4____ covalent bond. Pairs of valence electrons that are not shared between atoms are called ____5____. Sometimes two or three pairs of electrons may be shared to give ____6____ covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a ____7____ ____8____. It is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the ____9____.
- When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a ____10____.

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Rachel Sandford



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