

## Mole Concept

### Mole concept:

Mole is the counting unit for atoms, molecules and ions and is equal to  $6.022 \times 10^{23}$ . This number is called Avogadro's number and is denoted by  $N_A$ .

One mole is the amount of substance that contains as many particles (atoms, molecules or ions) as in 12g of C-12.

### Molar mass:

It is the mass of 1 mole of any substance expressed in grams.

In case of atoms it is equal to atomic mass expressed in grams.

In case of molecules it is equal to molecular mass expressed in grams.

1 Mole of  
particles  
(atoms,  
molecules or  
ions)

Contains  $6.022 \times 10^{23}$  particles  
(atoms,  
molecules or  
ions)

Mass is  
equal to  
molar mass  
(Atomic mass  
or molecular  
mass)

Volume is  
equal to  
molar  
volume =  
22.4 lit

For example,

|                              |   |      |          |
|------------------------------|---|------|----------|
| 1 mole of Oxygen gas         | Contains $6.022 \times 10^{23}$ Oxygen molecules.         | 32 g | 22.4 lit |
| 1 mole of Oxygen atoms       | Contains $6.022 \times 10^{23}$ Oxygen atoms.             | 16g  | 22.4 lit |
| 1 mole of Carbon dioxide gas | Contains $6.022 \times 10^{23}$ Carbon dioxide molecules. | 44g  | 22.4 lit |
| 1 mole of Ammonia gas        | Contains $6.022 \times 10^{23}$ Ammonia molecules.        | 17g  | 22.4 lit |

### Formula

1. No. of moles = given mass / molar mass ( $n = m/M$ )
2. No. of moles = given no. of particles / Avogadro no. ( $n = N/N_A$ )
3. Molar mass = given mass / no. of moles ( $M = m/n$ )
4. Mass of the substance = molar mass  $\times$  no. of moles ( $m = M \times n$ )
5. No. of particles = no. of moles  $\times$  Avogadro's no. (no. of particles =  $n \times N_A$ )  
= (Given Mass/ Molar mass)  $\times N_A$

# The Mole Concept Review Key

**Leo J. Malone, Theodore O. Dolter**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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**Daily Graphic** Elvis Aryeh, 2002-05-07 *Chemistry*<sup>3</sup>

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Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

*Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter **The Practice of Chemistry** Donald J. Wink, Sharon

Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt) 832KB

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