

CHAPTER 14: COVALENT BONDING: ORBITALS

PROBLEMS WITH LEWIS STRUCTURE

- Provides no indication of the **shapes** of the molecules
- O₂ would be predicted to be diamagnetic, but it is **paramagnetic**

Paramagnetic = has unpaired electrons.

- Many compounds cannot be represented by a single structure
- Resonance structures **differ only in position of electrons** not the nuclei.
- Result in notions of hypervalent compounds and "expansion of octet"
- Can't explain molecules such as B₂H₆ (electron deficient molecules)



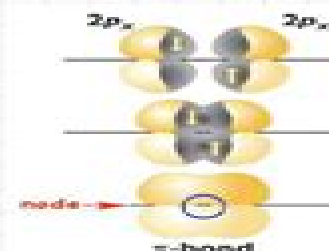
The Lewis structure of O₂ has an O=O double bond with **no unpaired electrons**, i.e. diamagnetic.

VALENCE-BOND THEORY "LOCALIZED ELECTRON MODEL"

- Covalent bond is created by the overlap of orbitals on two atoms

TYPES OF OVERLAP: σ and π Bonds

- Bonding in H₂**
 - Sausage shaped orbital, bond called **σ -orbital**
 - Called σ orbital because it has **no node** and is **cylindrically symmetrical**
 - Electron density is **concentrated between two nuclei**, but is still highest at the nuclei themselves.
- Bonding in N₂
 - σ orbital** is obtained by overlap of **two p_x-orbitals**
- Perpendicular p orbitals overlap in a side-by-side manner
- The resulting orbital has two lobes of electron density and is called a **π -bond**
- Called π -orbital because it **has a node** and, when viewed down the internuclear axis, it looks like a p-orbital
- π -bond has **no electron density between the nuclei along the bond axis**



University Chemistry Chapter Covalent Bonding Orbitals Study Guide

C. Barry Carter, M. Grant Norton



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undergraduate chemistry students and others taking courses of which chemistry is a significant part **Science Activities**, 1969 **Foundations of College Chemistry** Daniel Barker Murphy, Viateur Rousseau, 1969 *Physical Chemistry*

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Parsons, Gwen Pilling, Gareth Price, 2017 Chemistry3 establishes the fundamental principles of all three strands of chemistry organic inorganic and physical By building on what students have learned at school using carefully worded explanations annotated diagrams and worked examples it presents an approachable introduction to chemistry and its relevance to everyday life *High School Chemistry Teachers Magazine*, 1967 **Ceramic Materials** C. Barry Carter, M. Grant

Norton, 2013-01-04 Ceramic Materials Science and Engineering is an up to date treatment of ceramic science engineering and applications in a single comprehensive text Building on a foundation of crystal structures phase equilibria defects and the mechanical properties of ceramic materials students are shown how these materials are processed for a wide diversity of applications in today's society Concepts such as how and why ions move how ceramics interact with light and magnetic fields and how they respond to temperature changes are discussed in the context of their applications References to the art and history of ceramics are included throughout the text and a chapter is devoted to ceramics as gemstones This course tested text now includes expanded chapters on the role of ceramics in industry and their impact on the environment as well as a chapter devoted to applications of ceramic materials in clean energy technologies Also new are expanded sets of text specific homework problems and other resources for instructors The revised and updated Second Edition is further enhanced with color illustrations throughout the text

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