

MADONNA UNIVERSITY, NIGERIA
FIRST SEMESTER LECTURE TIME TABLE 2024/2025 ACADEMIC SESSION

200 LEVEL

COURSE CODE	COURSE TITLE	UNITS	DAYS	TIME	VENUE	LECTURER
-	Jagging Orientation	-	Mondays	7.00am - 9.00am	Field	-
MCB 201	Introduction to General Microbiology [MLS, NSC, PUH]	2	Mondays	10.00am - 12.00pm	NSC Auditorium	MCB Tutors
PIO 201	Introductory Physiology and Blood [ANA, PHY, MED]		Mondays	12.30pm - 2.30pm	NSC Auditorium	Physiology Tutors
			Mondays	2.30pm - 5.00pm		
ANA 201	Anatomy of Upper and Lower Limb	2	Tuesdays	8.00am - 10.00am	MLS Auditorium	Anatomy Tutors
BOH 201	General Biochemistry I [NSC, MLS, OPT, PUH, BOH, MCB]	2	Tuesdays	10.00am - 12.00pm	MLS Auditorium	Biochemistry Tutors
BOH 203	General Biochemistry Practical [PHARM, OPT, MLS, BOH, MB and IC]	1	Tuesdays	12.30pm - 2.30pm	BOH Labs	Biochemistry Tutors
ANA 203	General and Systemic Embryology	2	Tuesdays	2.30pm - 4.30pm	NSC Auditorium	Anatomy Tutors
			Wednesdays	8.00am - 10.00am		
PIO 201	Introductory Physiology and Blood [NSC, PUH, OPT]	2	Wednesdays	10.00am - 12.00pm	NSC Auditorium	Physiology Tutors
			Wednesdays	12.30pm - 2.00pm		
MUN-IMP 201	Igbo Metaphysics	1	Wednesdays	2.00pm - 4.00pm	Eden's Ark	Rev. Fr. Chinedu
-	Sport/Recreational Activities	-	Wednesdays	4.00pm - 6.00pm	Field	-
			Thursdays	8.00am - 10.00am		
PIO 201	Introductory Physiology and Blood [PHARM, MLS]	2	Thursdays	10.00am - 12.00pm	PHARM Building	Physiology Tutors
BOH 203	General Biochemistry Practical [NSC, ANA, PIO, MD]	1	Thursdays	1pm - 3.00pm	BOH Labs	Biochemistry Tutors
			Thursdays	3pm - 5.00pm		
STA 201	Statistics for Agriculture & Biological Sciences	3	Fridays	8.00am - 9.30am	Science Hall	
PIO 203	Physiology of Excitable Tissues	2	Fridays	10.00am - 12.00pm	MLS Auditorium	Physiology Tutors
ENT 211	Entrepreneurship and Innovation	2	Fridays	12.00pm - 2.00pm	Eden's Ark	Mwoko C. et al
MUN-THL 201	Fundamental Theology	1	Fridays	2.00pm - 3.30pm	Eden's Ark	Dr. Ezeobi

Note: 1. The courses highlighted in bold are fixed and cannot be changed. All departments **MUST** adjust accordingly and should not schedule any student activities during those times.

2. Departments should carefully review the available slots on the timetable to schedule their respective program-specific courses and any other course not captured.

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Terry C. Jones



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web jan 29 2020 the ionic form of this reaction is $2\text{Na}^{+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) + \text{Cu}^{2+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) \rightarrow 2\text{Na}^{+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{CuCl}_2(\text{s})$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web mar 11 2014 solution ionic equation $\text{H}^{+}(\text{aq}) + \text{Cl}^{-}(\text{aq}) + \text{Na}^{+}(\text{aq}) + \text{HCO}_3^{-}(\text{aq}) \rightarrow \text{Na}^{+}(\text{aq}) + \text{Cl}^{-}(\text{aq}) + \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ the spectator ions are Cl^{-} and Na^{+} net ionic equation $\text{H}^{+}(\text{aq}) + \text{HCO}_3^{-}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$ here is a video on spectator ions

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web the six NO_3^{-} ions and the six Na^{+} ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3\text{Ba}^{2+}(\text{aq}) + 2\text{PO}_4^{3-}(\text{aq}) \rightarrow \text{Ba}_3(\text{PO}_4)_2(\text{s})$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web video answer you know spectator ions are the ions present on the both sides of the equation indicate cancels out in the total ionic equations so you know in a case chemical reactions are described with chemical equations which symbolically present the starting materials end products

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web net ionic equations an equation for a reaction in solution showing only those particles that are directly involved in the chemical change spectator ions an ion that is not directly involved in a chemical reaction an ion that does not change oxidation number or composition during a reaction products

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^{+}(\text{aq}) + \text{OH}^{-}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$

2 o l the spectator ions in koh aq and hno₃ aq are k⁺ aq and no₃⁻ aq how to identify and list the

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web apr 6 2022 spectator ion a spectator ion is an ion that exists as a reactant and a product in a chemical equation a spectator ion can therefore be observed in the reaction of aqueous solutions of sodium carbonate and copper ii sulfate but does not affect the equilibrium what are spectator ions

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web the ag⁺ aq and cl⁻ aq ions become agcl s but the na⁺ aq ions and the no₃⁻ aq ions stay as na⁺ aq ions and no₃⁻ aq ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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web the new format requires students to write balanced chemical equations showing only the reacting substances excluding spectator ions for three different sets of reactants and to answer a short question requiring no calculator about each reaction the questions are intended to test the students understanding of the meaning of the reactions

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web jun 2 2023 the spectator ions are cut out from a reaction hence the net ionic equation for the above reaction would be

$2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + 2\text{Cl}^- \text{aq} + 2\text{S} \rightarrow 2\text{Cl}^- \text{aq} + \text{Cu}^{2+} \text{aq} + 2\text{Cl}^- \text{aq} + 2\text{S}$ these spectator ions are not actively involved in the reaction the existence of spectator ions affects the Debye length