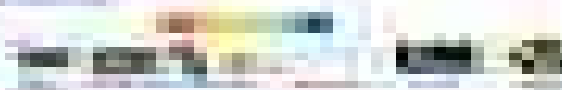


Topics

- 1. **Thermodynamics** - study of energy and its transformations in systems and matter



- 2. **The First Law of Thermodynamics** - Energy is conserved in a closed system. $\Delta U = Q - W$ (change in internal energy = heat added - work done).



- 3. **Heat capacity** - amount of heat required to raise the temperature of a substance by 1 degree Celsius. $C = \frac{Q}{\Delta T}$.



- 4. **Entropy** - measure of disorder or randomness in a system. $\Delta S = \frac{Q}{T}$ (change in entropy = heat added / temperature).



- 5. **The Second Law of Thermodynamics** - Entropy of a closed system always increases over time.



Topics

- 6. **Heat engines** - devices that convert heat into work. Efficiency is the ratio of work done to heat added.

- 7. **Heat pumps** - devices that transfer heat from a cold reservoir to a hot reservoir.



- 8. **Refrigerators** - devices that transfer heat from a cold reservoir to a hot reservoir. Coefficient of performance (COP) is the ratio of heat removed to work done.



- 9. **Thermodynamic potentials** - functions that describe the state of a system. Examples include internal energy (U), enthalpy (H), and Gibbs free energy (G).

- (a) Internal energy (U)
- (b) Enthalpy (H)
- (c) Gibbs free energy (G)

- 10. **Statistical mechanics** - study of the microscopic behavior of particles and how it relates to macroscopic thermodynamic properties.



- (a) Microscopic behavior of particles
- (b) Macroscopic thermodynamic properties
- (c) Relationship between the two

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Brendan G. Carr



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