

Reactions of Amines and Epoxides

Preparation of Amines

Amination of alcohols: Polyalkylation is a serious problem for tertiary hydroxyl group



A summary of Ring-Opening Reactions of Epoxides with Strong and Weak Nucleophiles



Bond Types, Strength, Length, Polarity and Dipole Moment

Two different bond types

There are two types of bonding: **ionic bonding** and **covalent bonding**

- **Ionic bonds** are formed by the transfer of electrons from one element to another (ionization potential) and electron affinity. Examples:



- **Covalent bonds** are formed by sharing of electrons between two atoms.



The Relative Electronegativity of an



Bonded and Unbonded electron



Bonding electrons



There are two main categories

Definition of electron pair



Sigma and pi bonds

Covalent bonds are classified in **sigma** and **pi** bonds

- **Sigma bonds** are formed by head-on (frontal) overlap of orbitals.
- **pi bonds** are formed by side-on overlap of orbitals.
- **Any single bond** is a **sigma bond**.
- **Any multiple bond** consists of one **sigma** and **n** **pi** bonds.



pi bonds are stronger than sigma bonds.

Polar and Nonpolar Covalent Bonds

Depending on the difference in the electronegativity values, covalent bonds can be **polar** and **nonpolar**.



- If the difference in electronegativity is **less than 0.5**, the electrons are shared equally (nonpolar bond). For example, **C-C** and **C-H** bonds.

- If the difference in electronegativity is **between 0.5 and 1.7**, we have a **polar covalent bond**. More electronegative atom attracts the electron pair more. For example, **C-O**, **C-F** and other halogenes.

- A difference of **1.7 or higher** can be high that the electrons are not longer shared and a **ionic bond** is formed.

Dipole Moment

The polarity of a bond is described by the **dipole moment** (μ).



- The **dipole moment** is always a vector pointing from the less electronegative to toward the more electronegative atom.
- The more electronegative atom has a partial negative charge (δ^-) and the less electronegative atom has a partial positive charge (δ^+).
- **δ^- (delta minus)** indicates an electron deficient atom.
- **δ^+ (delta plus)** indicates an electron rich atom.

Molecular Dipole

To determine the **molecular polarity**, the vector sum of the dipoles is used.

The geometry of the molecule is important. Symmetrical molecules are nonpolar even if polar bonds are present. A good example is carbon dioxide.



Consider the water as the other polarizing point in the water case.

Bond Length and Bond Strength

The longer the bond, the stronger it is.

| Bond | H-H | H-Cl | H-Br | H-I |
|----------|------------|------------|------------|------------|
| Length | 109 pm | 127 pm | 141 pm | 154 pm |
| Strength | 436 kJ/mol | 431 kJ/mol | 366 kJ/mol | 298 kJ/mol |

Double bonds are stronger than single bonds and a triple bond is stronger than a double bond.

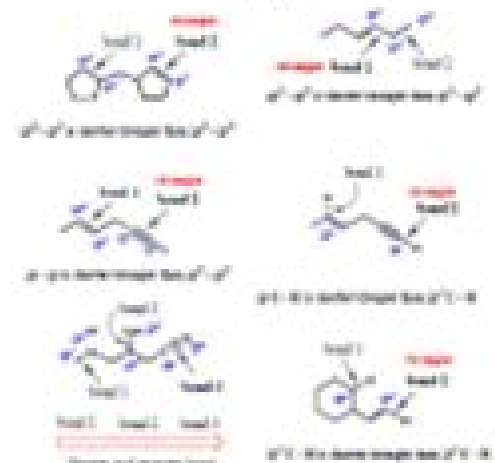


The longer the bond, the weaker it is.

| Average values of C-C and C-H bonds depending on hybridization | | | |
|--|-----------------|-----------------|-------------|
| C-C Bond | sp^3-sp^3 C-C | sp^2-sp^2 C-C | $sp-sp$ C-C |
| Hybridization | sp^3-sp^3 | sp^2-sp^2 | $sp-sp$ |
| Length | 153 pm | 134 pm | 120 pm |
| Strength | 376 kJ/mol | 611 kJ/mol | 837 kJ/mol |
| Hybridization | sp^3-sp^2 | sp^2-sp | $sp-sp$ |
| C-H Bond | 109 pm | 108 pm | 106 pm |
| Length | 153 pm | 134 pm | 120 pm |
| Strength | 376 kJ/mol | 611 kJ/mol | 837 kJ/mol |

Higher s character = stronger and shorter bond

Examples



Study Guide For Organic Chemistry

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