Chemistry Finals Study Guide

Chapter 1 & 2 - Review of Basic Concepts

- Describe the 3 states of matter. The three states of matter are solid, liquid, and gas.
 - Solid: When atoms in a material are moving slow that they stick together and barely move, creating a hard object classified as a solid
 - Liquid: When atoms are moving slower than a gas state but faster than the solid state so they flow and are able to take the shape of their container. They have more energy than solid state atoms.
 - Gas: When atoms are moving faster than the other two states of matter (solid & liquid) that they are no longer pulled towards each other and roam freely without boundaries if not kept in a container.
- Describe the difference between elements and compounds, mixtures.
 - Elements: materials that are only made out of the same type of atoms and cannot be changed into simpler substances by a chemical change (ex: O₃)
 - Compounds: materials that are a mix of two or more element in a fixed proportion
 - Mixtures: a mixture of 2 or more pure substances.
 - Heterogeneous Mixture: can often be separated by filtration
 - Homogenous Mixture: can be separated by distillation, crystallization, and chromatography
- Convert between Celsius and Kelvin.
 - Celsius to Kelvin: C = (°K 273.15)
 - Kelvin to Celsius: K = (°C + 273.15)
- Round measured and calculated values to the correct number of significant digits.
 - Significant Digit Rules: [CHEM] Final 1 Steps & Processes
- Use dimensional analysis to solve multi step equations, showing how units cancel
 [CHEM] Final 1 Steps & Processes
- What is the difference between precision and accuracy?
 - Precision: Acquiring the same result again and again
 - Accuracy: A result that is close to the accepted value/actual answer

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