

# Gas Laws Practice

Name: \_\_\_\_\_

Date: \_\_\_\_\_

Hour: \_\_\_\_\_

**IMPORTANT:** whenever you use temperature, it must be in degree Kelvin (K), so remember the equation:  $K = ^\circ C + 273$

1. a) convert  $39\text{ }^\circ\text{C}$  to K.                      b) convert  $127\text{ K}$  to  $^\circ\text{C}$ .
2. A gas has an initial volume of  $2.75\text{ L}$  at a temperature of  $285\text{ K}$ . If the temperature changes to  $380\text{ K}$ , what is the new volume of the gas if the pressure is unchanged?
3. Gas can often be cooled by compressing it while keeping the pressure constant. If I have  $45.0\text{ L}$  of gas at room temperature ( $22^\circ\text{C}$ ) and I compress it so that the final volume is  $0.50\text{ L}$ , what is the final temperature of the gas if the pressure is constant?
4. The volume of a gas is  $2.5\text{ L}$  when the pressure is at standard pressure ( $101.325\text{ kPa}$ ). What is the volume of the gas if the pressure decreases to  $85\text{ kPa}$  and the temperature remains unchanged?
5. A  $5.0\text{ L}$  container of gas experiences a temperature change so that the final temperature is 4 times the initial temperature. What is the size of the container after the temperature change? (Assume constant pressure.)
6. At  $45^\circ\text{C}$  the volume of a certain gas is  $27.5\text{ L}$  and the pressure is  $210\text{ kPa}$ . What is the volume of the gas at standard temperature ( $273\text{ K}$ ) and  $310\text{ kPa}$  of pressure?
7. The pressure of a sample of gas was  $97.8\text{ kPa}$  and the volume of the gas was  $3.75\text{ L}$ . If the gas occupied a container with a volume of  $8.00\text{ L}$ , what would the pressure in the container be?
8. Isothermal expansion refers to allowing a gas to expand while keeping the temperature constant. This is one means to simulate a vacuum. If a gas originally at  $97\text{ kPa}$  is allowed to expand from  $0.25\text{ L}$  to  $182\text{ L}$ , what is the pressure of the gas?
9. A gas is initially at a pressure of  $225\text{ kPa}$  and a temperature of  $245\text{ K}$  in a container that is  $4.5\text{ L}$ . If the gas is compressed to a volume of  $2.1\text{ L}$  and the temperature changes to  $275\text{ K}$ , what is the new pressure?

# Skill Practice 3gas Laws Practice Answers

**L Reisser**



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