

1. Write your answer.

Math Skills

Determining pH

Below are eight sample problems and solutions using the equations provided and as separate sheets of paper. Write your answers to the corresponding problem.

Problem 1a

For the weak acid H_2A with $K_{\text{a1}} = 1.0 \times 10^{-4}$ calculate the pH of a 0.10 M solution of H_2A .

Solution 1a

Step 1: Write the given and unknown values.

Given: concentration of H_2A is 0.10 M; $K_{\text{a1}} = 1.0 \times 10^{-4}$

Unknown: pH

Step 2: Consider the weak acid dissociation of H_2A in the water.

concentration of H_2A is solution is 0.10 M; $K_{\text{a1}} = 1.0 \times 10^{-4}$

H_2A is a weak acid, so the concentration of H_2A is not significantly changed in the solution. In equation the concentration of H_2A

concentration of H_2A is 0.10 M; $K_{\text{a1}} = 1.0 \times 10^{-4}$ is 0.10 M; $K_{\text{a1}} = 1.0 \times 10^{-4}$

Step 3: Calculate the pH of the solution.

$\text{pH} = -\log[\text{H}^+]$ is the negative of the power of ten used to describe the concentration of H^+ ions.

concentration of H^+ is 0.001 M; $K_{\text{a1}} = 1.0 \times 10^{-4}$

$\text{pH} = -\log(0.001)$

Problem 1b

1. A weak acid, H_2A , has a dissociation constant of 1.0×10^{-4} . If the solution is 0.10 M H_2A , what is the pH of the solution?

2. A weak acid, H_2A , has a dissociation constant of 1.0×10^{-4} . If the solution is 0.10 M H_2A , what is the pH of the solution?

Skills Math Skills Determining Ph

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