

Solution and Answer Guide

Solution and Answer Guide: Skoog et al., *Fundamentals of Analytical Chemistry*, 10e,
© 2022, 978-0-357-45039-0, Chapter 11: Titrations in Analytical Chemistry

Chapter 11

11-1. Define

Answers:

***(a)** millimole.

The millimole is the amount of an elementary species, such as an atom, an ion, a molecule, or an electron. A millimole contains

$$6.02 \times 10^{23} \frac{\text{particles}}{\text{mol}} \times \frac{\text{mmol}}{1000 \text{ mol}} = 6.02 \times 10^{20} \frac{\text{particles}}{\text{mmol}}$$

(b) titration.

A titration measures the quantity of a reagent of known concentration required to react with a measured quantity of sample of an unknown concentration. The concentration of the sample is then determined from the quantities of reagent and sample, the concentration of the reagent, and the stoichiometry of the reaction.

***(c)** stoichiometric ratio.

The stoichiometric ratio is the molar ratio of two chemical species that appear in a balanced chemical equation.

(d) titration error.

Titration error is the error encountered in titrimetry that arises from the difference between the amount of reagent required to give a detectable end point and the theoretical amount for reaching the equivalence point.

11-2. Write two equations that—along with the stoichiometric factor—form the basis for the calculations of volumetric titrations.

Solution:

$$\text{amount A (mmol)} = \text{volume (mL)} \times c_A \text{ (mmol A/mL)}$$

$$\text{amount A (mmol)} = \frac{\text{mass A (g)}}{\text{millimolar mass A (g/mmol)}}$$

*Answers and solutions are provided in the Student Solutions Manual for questions and problems marked with an asterisk.

Skoog Spectroscopy Answer Guide

RJ Shavelson



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