

SOLUBILITY CURVES

Name _____

Answer the following questions based on the solubility curve below.

- Which salt is least soluble in water at 20° C? KClO₃
- How many grams of potassium chloride can be dissolved in 200 g of water at 80° C?
100 g
- At 40° C, how much potassium nitrate can be dissolved in 300 g of water? 123 g
- Which salt shows the least change in solubility from 0° - 100° C?
NaCl
- At 30° C, 90 g of sodium nitrate is dissolved in 100 g of water. Is this solution saturated, unsaturated or supersaturated?
unsaturated
- A saturated solution of potassium chlorate is formed from one hundred grams of water. If the saturated solution is cooled from 80° C to 50° C, how many grams of precipitate are formed? 40 - 20 g = 20 g
- What compound shows a decrease in solubility from 0° to 100° C? NH₃
- Which salt is most soluble at 10° C? KI
- Which salt is least soluble at 50° C? KClO₃
- Which salt is least soluble at 90° C? NH₃



Solubility Curve Practice Problems 1 Answer Key

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Chemistry Calculations for Beginners John Obimakinde, Samuel Obimakinde, Ebenezer Obimakinde, Fredrick Akinbolade, 2025-05-30 With decades of combined experience as science teachers at both school and undergraduate levels the authors have recognised that one of the greatest challenges faced by students studying chemistry is grasping the complexity of the numerous numerical problems found in most parts of the subject This text is crafted to provide a clear and accessible pathway to overcoming this challenge by assisting students especially novices or those with minimal knowledge of the subject in performing chemistry calculations The content covers fundamental calculations crucial to understanding the principles of chemistry making it an invaluable tool for students aiming to excel in their studies Key features Designed with a student friendly approach including detailed explanation of chemical concepts underlying each type of calculation step by step explanations alternative methods for solving problems numerous practice exercises answers to practice exercises and appendices The book is tailored to suit various curricula ensuring relevance for a diverse audience Encompasses a wide range of calculations offering students a thorough understanding of essential chemistry concepts Serves as an excellent resource for exam preparation and equips students with skills applicable to future scientific endeavours Employs straightforward language to ensure ease of understanding for beginners Uses IUPAC conventions underscoring the universal nature of chemistry

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