

Solubility Rules Worksheet

1. Classify each substance as being soluble or insoluble in water.

- | | |
|--------------------------------|---------------------------------|
| a. Potassium bromide – _____ | f. Silver iodide - _____ |
| b. Lead (II) carbonate – _____ | g. Cadmium (II) sulfide - _____ |
| c. Barium sulfate – _____ | h. Zinc carbonate - _____ |
| d. Zinc hydroxide – _____ | i. Silver acetate - _____ |
| e. Sodium acetate – _____ | j. Copper (II) sulfide – _____ |

2. Identify the two new compounds which form if the solutions were mixed. Circle any compounds that would precipitate from the solution (are insoluble).

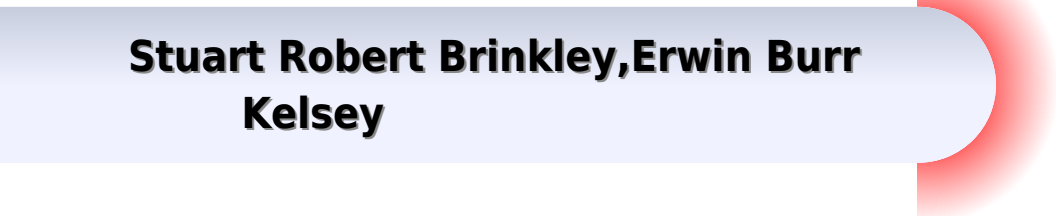
	KBr	Na_2CO_3	CaS	NH_4OH
AgNO_3				
BaCl_2				
$\text{Al}(\text{NO}_3)_3$				
CuSO_4				

3. Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formu	Name	Solubility
$\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$		
	Barium hydroxide	
FeCO_3		
NaOH		
	Rubidium nitrate	
MgSO_4		
	Zinc chloride	
	Zinc hydroxide	

Solubility Rules Activity Answers

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Kelsey**



Solubility Rules Activity Answers:

ChemDiscovery Lab Manual Olga I. Agapova,2002 *You Want Me to Teach What?* Norman Joseph LaFave,2012

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Aquatic Chemistry Werner Stumm, James J. Morgan, 2013-09-23 The authoritative introduction to natural water chemistry THIRD EDITION Now in its updated and expanded Third Edition Aquatic Chemistry remains the classic resource on the essential concepts of natural water chemistry Designed for both self study and classroom use this book builds a solid foundation in the general principles of natural water chemistry and then proceeds to a thorough treatment of more advanced topics Key principles are illustrated with a wide range of quantitative models examples and problem solving methods Major subjects covered include Chemical Thermodynamics Solid Solution Interface and Kinetics Trace Metals Acids and Bases Kinetics of Redox Processes Dissolved Carbon Dioxide Photochemical Processes Atmosphere Water Interactions Kinetics at the Solid Water Metal Ions in Aqueous Solution Interface Precipitation and Dissolution Particle Particle Interaction Oxidation and Reduction Regulation of the Chemical Equilibria and Microbial Mediation Composition of Natural Waters

Chemistry Charles H. Corwin, 1994 The book focuses on the concepts of chemistry and the applications that maintain and generate motivation for the subject of chemistry

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