

Solubility Rules

When salt solutions are mixed, two things can happen: 1) no reaction (all ions remain soluble), or 2) a solid forms (called a precipitate). You will examine which salts (positive + negative ions) are soluble and which are insoluble.

Safety: Wear goggles, wash your hands after lab.

Hypothesis: Look at the grid below. Looking back at your solubility rules sheet, which of these reactions do you expect to form a precipitate (solid)? _____

Procedure

1. Obtain from teacher: 1 well plate and aqueous solutions to be used in investigations.
2. Place your well plate on the lab bench.
3. Using the dropper bottles at your lab station, carefully add one drop (from a height of 1-2 cm) to each well for that ion. E.g. for Br^- a drop goes in each box under Br^- ; for Ag^+ a drop goes in each box left of Ag^+ . (be careful to note the label on each bottle). It is critical to add drops from a height of 1-2 cm. If you touch the tip of the dropper bottle to another solution, the contents of that bottle can become contaminated, ruining the results of your group and other students. Continue this procedure until all ions have been added together. At the end of this step you should have, in each box, a mixture of two drops (1 positive ion + 1 negative ion).
4. Look in the boxes: some will contain a precipitate (i.e. solid), others will be clear (indicating the salt is soluble).
5. On the grid below, place an "X" in every box where a precipitate has formed.
6. Rinse off the reaction grid in the sink. Rinse out the sink. Wipe the grid clean using a damp paper towel. Dry the grid. Return all equipment. Wipe off your lab bench. Wash your hands thoroughly.

	AgNO_3 (Ag^+)	$\text{Pb}(\text{NO}_3)_2$ (Pb^{+2})	CaCl_2 (Ca^{+2})	CuSO_4 (Cu^{+2})	$\text{Fe}(\text{NO}_3)_3$ (Fe^{+3})	K_2CrO_4 (K^+)
Na_3PO_4 (PO_4^{-3})	a	d	g	j	m	p
NaOH (OH^{-1})	b	e	h	k	n	q
Na_2SO_4 (SO_4^{-2})	c	f	i	l	o	r

Solubility Rules And The Mystery Solutions Lab Answers

Mark Steyvers



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