

CHAPTER

02

SOLUTION

1. **Solution** : A homogeneous mixture of two or more substances is known as solution.
2. **Solute** : The substance present in smaller amount in a solution is called solute.
3. **Solvent** : The substance present in larger amount in a solution is called solvent.
4. **Binary solution** : Those solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
5. **Ternary solutions** : Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
6. **Aqueous solution** : When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
7. **Non-aqueous solution** : When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
8. **Weight Fraction** = $\frac{\text{weight of solute}}{\text{weight of solution}}$
9. **Weight percentage**

$$= \frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage
10. **Volume percentage**

$$= \frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
11. **Volume Fraction** = $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
12. **Molarity (M)**

$$= \frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1}{V(\text{lit})}$$
13. **Molarity** = $\frac{\text{No of moles of solute}}{\text{Volume of solution in ml} \times \frac{1000}{1000}}$
14. **Molarity** = $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1000}{V(\text{ml})}$$
15. **Weight of solute** = $\frac{M \times GMW \times V \text{ in ml}}{1000}$
16. **Molarity for dilute solutions** : $M_1 V_1 = M_2 V_2$
17. **For neutralisation reactions** : $\frac{M_1 V_1}{n_1} = \frac{M_2 V_2}{n_2}$
18. **For mixed solution molarity**

$$= \frac{M_1 V_1 + M_2 V_2 + \dots}{V_1 + V_2 + \dots}$$

Solution For All

Roman Wölfel



Solution For All:

The book delves into Solution For All. Solution For All is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Solution For All, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Solution For All
 - Chapter 2: Essential Elements of Solution For All
 - Chapter 3: Solution For All in Everyday Life
 - Chapter 4: Solution For All in Specific Contexts
 - Chapter 5: Conclusion
2. In chapter 1, the author will provide an overview of Solution For All. This chapter will explore what Solution For All is, why Solution For All is vital, and how to effectively learn about Solution For All.
3. In chapter 2, this book will delve into the foundational concepts of Solution For All. The second chapter will elucidate the essential principles that must be understood to grasp Solution For All in its entirety.
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5. In chapter 4, the author will scrutinize the relevance of Solution For All in specific contexts. The fourth chapter will explore how Solution For All is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, the author will draw a conclusion about Solution For All. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Solution For All.

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