

CHAPTER 2

ATOMIC STRUCTURE AND INTERATOMIC BONDING

PROBLEM SOLUTIONS

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Fundamental Concepts

Electrons in Atoms

2.1 Cite the difference between atomic mass and atomic weight.

Solution

Atomic mass is the mass of an individual atom, whereas atomic weight is the average (weighted) of the atomic masses of an atom's naturally occurring isotopes.

2.2 Silicon has three naturally occurring isotopes: 92.23% of ^{28}Si , with an atomic weight of 27.9769 amu, 4.68% of ^{29}Si , with an atomic weight of 28.9765 amu, and 3.09% of ^{30}Si , with an atomic weight of 29.9738 amu. On the basis of these data, confirm that the average atomic weight of Si is 28.0854 amu.

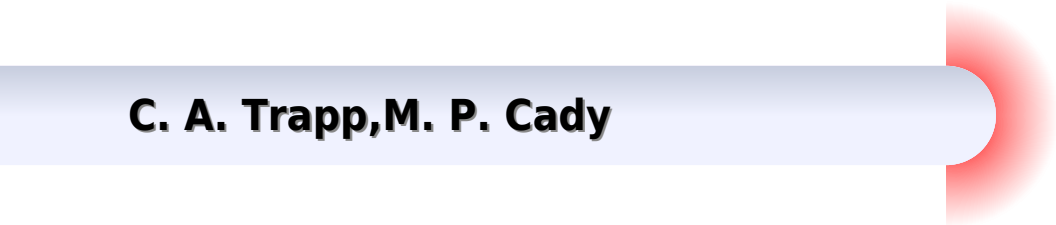
Solution

The average atomic weight of silicon (A_{Si}) is computed by adding fraction-of-occurrence/atomic weight products for the three isotopes—i.e., using Equation 2.2. (Remember: fraction of occurrence is equal to the percent of occurrence divided by 100.) Thus

$$\begin{aligned} A_{\text{Si}} &= f_{28} A_{28} + f_{29} A_{29} + f_{30} A_{30} \\ &= (0.9223)(27.9769) + (0.0468)(28.9765) + (0.0309)(29.9738) = 28.0854 \end{aligned}$$

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C. A. Trapp, M. P. Cady



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1. Understanding the eBook Solution Manual Atomic Bonding
 - The Rise of Digital Reading Solution Manual Atomic Bonding
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual Atomic Bonding
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual Atomic Bonding
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Manual Atomic Bonding

- Personalized Recommendations
 - Solution Manual Atomic Bonding User Reviews and Ratings
 - Solution Manual Atomic Bonding and Bestseller Lists
5. Accessing Solution Manual Atomic Bonding Free and Paid eBooks
 - Solution Manual Atomic Bonding Public Domain eBooks
 - Solution Manual Atomic Bonding eBook Subscription Services
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 - ePub, PDF, MOBI, and More
 - Solution Manual Atomic Bonding Compatibility with Devices
 - Solution Manual Atomic Bonding Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solution Manual Atomic Bonding
 - Highlighting and Note-Taking Solution Manual Atomic Bonding
 - Interactive Elements Solution Manual Atomic Bonding
 8. Staying Engaged with Solution Manual Atomic Bonding
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual Atomic Bonding
 9. Balancing eBooks and Physical Books Solution Manual Atomic Bonding
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual Atomic Bonding
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Solution Manual Atomic Bonding
 - Setting Reading Goals Solution Manual Atomic Bonding
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Solution Manual Atomic Bonding

- Fact-Checking eBook Content of Solution Manual Atomic Bonding
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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