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Masterton Hurley ISBN 130507937X 9781305079373**

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2 : ATOMS, MOLECULES AND IONS

LECTURE NOTES

Students find this material relatively easy to assimilate; it's almost entirely qualitative. On the other hand, there's a lot of memorizing (sorry, learning) to do. This chapter is coverable in two lectures.

Some general observations:

- ✓ Material in Sections 2.1–2.3 is generally well covered in high-school chemistry courses; no need to dwell on it.
- ✓ Students need to know the molecular formulas of the elements (Figure 2.13), the charges of ions with noble-gas structures and the names and formulas of the common polyatomic ions (Table 2.2). The charges of transition-metal ions will be covered later, in Chapter 4.
- ✓ Naming compounds requires students to distinguish between ionic and molecular substances. It helps to point out that binary molecular compounds are composed of two nonmetals. Almost all ionic compounds contain a metal cation combined with a nonmetal anion or negatively charged polyatomic ion. The flow charts shown in Figures 2.18 and 2.19 should help visual learners.
- ✓ The periodic table will be discussed in greater detail later in the text (Chapter 6).

Lecture 1

I. Atomic Theory

A. Elements

Postulates: Elements consist of tiny particles called atoms, which retain their identity in reactions. In a compound, atoms of two or more elements combine in a fixed ratio of small whole numbers (e.g., 1:1, 2:1, etc.).

B. Components

	relative mass	relative charge	location
proton	1	+1	nucleus
neutron	1	0	nucleus
electron	0.0005	1	outside

Solution Manual Chemistry Principles And Reactions

William L Masterton, Cecile N. Hurley



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Student Solutions Manual for Masterton/Hurley's Chemistry: Principles and Reactions, 8th William L. Masterton, Cecile N. Hurley, 2015-07-13 Improve your performance at exam time with this manual's complete solutions to the even numbered end of chapter Questions and Problems answered in Appendix 5 including the Challenge Problems The authors include references to textbook sections and tables to help guide you through the problem solving techniques employed by the authors Student Solutions Manual, Chemistry, Principles and Reactions, Seventh Edition, 2011

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Chemistry William L. Masterton, Cecile N. Hurley, 2011-08-26 The Student Solutions Manual contains complete solutions to all end of chapter Questions and Problems answered in Appendix 6 including the Challenge Problems The author includes references to textbook sections and tables to help guide students to use the problem solving techniques employed by authors

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