

CHAPTER 18 ACID-BASE EQUILIBRIA

FOLLOW-UP PROBLEMS

- 18.1A **Plan:** Examine the formulas and classify each as an acid or base. Strong acids are the hydrohalic acids HCl, HBr, and HI, and oxoacids in which the number of O atoms exceeds the number of ionizable protons by at least two. Other acids are weak acids. Strong bases are soluble oxides or hydroxides of the Group 1A(1) metals and Ca, Sr, and Ba in Group 2A(2). Other bases are weak bases.

Solution:

- a) Chloric acid, HClO_3 , is the stronger acid because acid strength increases as the number of O atoms in the acid increases.
b) Hydrochloric acid, HCl , is one of the strong hydrohalic acids whereas acetic acid, CH_3COOH , is a weak carboxylic acid.
c) Sodium hydroxide, NaOH , is a strong base because Na is a Group 1A(1) metal. Methylamine, CH_3NH_2 , is an organic amine and, therefore, a weak base.

- 18.1B **Plan:** Examine the formulas and classify each as a strong acid, weak acid, strong base, or weak base. Strong acids are the hydrohalic acids HCl, HBr, and HI, and oxoacids in which the number of O atoms exceeds the number of ionizable protons by at least two. Other acids are weak acids. Strong bases are soluble oxides or hydroxides of the Group 1A(1) metals and Ca, Sr, and Ba in Group 2A(2). Other bases are weak bases.

Solution:

- a) $(\text{CH}_3)_3\text{N}$ is a **weak base**. It contains a nitrogen atom with a lone pair of electrons, which classifies it as a base; however, it is not one of the strong bases.
b) Hydroiodic acid, HI , is a **strong acid** (one of the strong acids listed above).
c) HBrO is a **weak acid**. It has an ionizable hydrogen, which makes it an acid. Specifically, it is an oxoacid, in which a polyatomic ion is the anion. In the case of this oxoacid, there is only one O atom for each ionizable hydrogen, so this is a weak acid. The reaction for the dissociation of this weak acid is:
 $\text{HBrO}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{BrO}^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$. The corresponding equilibrium expression is:

$$K_a = \frac{[\text{BrO}^-][\text{H}_3\text{O}^+]}{[\text{HBrO}]}$$

- d) $\text{Ca}(\text{OH})_2$ is a **strong base** (one of the strong bases listed above).

- 18.2A **Plan:** The product of $[\text{H}_3\text{O}^+]$ and $[\text{OH}^-]$ remains constant at 25°C because the value of K_w is constant at a given temperature. Use $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ to solve for $[\text{H}_3\text{O}^+]$.

Solution:

Calculating $[\text{H}_3\text{O}^+]$:

$$[\text{H}_3\text{O}^+] = \frac{K_w}{[\text{OH}^-]} = \frac{1.0 \times 10^{-14}}{6.7 \times 10^{-2}} = 1.4925 \times 10^{-13} = \mathbf{1.5 \times 10^{-13} \text{ M}}$$

Since $[\text{OH}^-] > [\text{H}_3\text{O}^+]$, the solution is **basic**.

- 18.2B **Plan:** The product of $[\text{H}_3\text{O}^+]$ and $[\text{OH}^-]$ remains constant at 25°C because the value of K_w is constant at a given temperature. Use $K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ to solve for $[\text{H}_3\text{O}^+]$.

Solution:

Calculating $[\text{OH}^-]$:

$$[\text{OH}^-] = \frac{K_w}{[\text{H}_3\text{O}^+]} = \frac{1.0 \times 10^{-14}}{1.8 \times 10^{-10}} = 5.5555 \times 10^{-5} = \mathbf{5.6 \times 10^{-5} \text{ M}}$$

Since $[\text{OH}^-] > [\text{H}_3\text{O}^+]$, the solution is **basic**.

- 18.3A **Plan:** NaOH is a strong base that dissociates completely in water. Subtract pH from 14.00 to find the pOH, and calculate inverse logs of pH and pOH to find $[\text{H}_3\text{O}^+]$ and $[\text{OH}^-]$, respectively.

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Table of Contents Solution Manual For Ch 18 Chemical Equilibrium

1. Understanding the eBook Solution Manual For Ch 18 Chemical Equilibrium
 - The Rise of Digital Reading Solution Manual For Ch 18 Chemical Equilibrium
 - Advantages of eBooks Over Traditional Books
2. Identifying Solution Manual For Ch 18 Chemical Equilibrium
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solution Manual For Ch 18 Chemical Equilibrium
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solution Manual For Ch 18 Chemical Equilibrium

- Personalized Recommendations
 - Solution Manual For Ch 18 Chemical Equilibrium User Reviews and Ratings
 - Solution Manual For Ch 18 Chemical Equilibrium and Bestseller Lists
5. Accessing Solution Manual For Ch 18 Chemical Equilibrium Free and Paid eBooks
- Solution Manual For Ch 18 Chemical Equilibrium Public Domain eBooks
 - Solution Manual For Ch 18 Chemical Equilibrium eBook Subscription Services
 - Solution Manual For Ch 18 Chemical Equilibrium Budget-Friendly Options
6. Navigating Solution Manual For Ch 18 Chemical Equilibrium eBook Formats
- ePub, PDF, MOBI, and More
 - Solution Manual For Ch 18 Chemical Equilibrium Compatibility with Devices
 - Solution Manual For Ch 18 Chemical Equilibrium Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Solution Manual For Ch 18 Chemical Equilibrium
 - Highlighting and Note-Taking Solution Manual For Ch 18 Chemical Equilibrium
 - Interactive Elements Solution Manual For Ch 18 Chemical Equilibrium
8. Staying Engaged with Solution Manual For Ch 18 Chemical Equilibrium
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solution Manual For Ch 18 Chemical Equilibrium
9. Balancing eBooks and Physical Books Solution Manual For Ch 18 Chemical Equilibrium
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solution Manual For Ch 18 Chemical Equilibrium
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solution Manual For Ch 18 Chemical Equilibrium
- Setting Reading Goals Solution Manual For Ch 18 Chemical Equilibrium
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solution Manual For Ch 18 Chemical Equilibrium

- Fact-Checking eBook Content of Solution Manual For Ch 18 Chemical Equilibrium
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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