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## Chapter 2 – Atoms, Ions, and the Periodic Table

- 2.1 (a) neutron; (b) law of conservation of mass; (c) proton; (d) main-group element; (e) relative atomic mass; (f) mass number; (g) isotope; (h) cation; (i) subatomic particle; (j) alkali metal; (k) periodic table
- 2.2 (a) transition element; (b) law of definite proportions; (c) electron; (d) anion; (e) diatomic molecule; (f) noble gas; (g) period; (h) atomic number; (i) atomic mass unit (amu); (j) group or family; (k) alkaline earth metal; (l) nucleus
- 2.3 Dulton used the laws of conservation of mass (Lavoisier) and definite proportions (Proust). Dulton essentially reasoned that, because pure substances were always composed of elements in some fixed ratios, matter must be composed of discrete units (atoms).
- 2.4 The scanning tunneling microscope (STM) is used to image atoms on the surface of a material (Figure 2.4).
- 2.5 The second postulate of Dalton's atomic theory, listed in section 2.1, states that atoms of different elements differ in their atomic masses and chemical properties.
- 2.6 The third and fourth postulates of Dulton's atomic theory, listed in section 2.1, can be used to explain conservation of mass. In essence, because atoms only rearrange and no new matter can be formed during chemical reactions, mass must be conserved. If elements could be changed into other elements during chemical reactions (as the alchemists were trying to do), then masses of atoms would change during reactions and mass would not be conserved.
- 2.7 Compounds contain discrete numbers of atoms of each element that form them. Because all the atoms of an element have the same relative atomic mass, the mass ratio of the elements in a compound is always the same (law of definite proportions).
- 2.8 The fourth postulate of Dalton's atomic theory (section 2.1) states that compounds are formed when elements combine in simple, whole-number ratios. In addition, for any pure substance that ratio must be fixed.
- 2.9 No. On the left side of the diagram there are four white atoms and two blue atoms (represented as diatomic molecules). On the right, however, there are six white atoms. Since the numbers of each atom are not conserved, mass is not conserved. For mass to be conserved, another white molecule is needed on the left side of the diagram.
- 2.10 Yes. There are four white atoms and four green atoms on the left side. On the right, the atoms have recombined to form new substances, but the numbers of each atom are the same. This means mass has been conserved.

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**Rachel Sandford** 



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