



CE

#### GENERAL INDICATIONS

POWER → COMM. FAULT → COMM. FAULT → SYSTEM FAULT →  
SUPPLY FAULT → BATTERY FAULT → BATTERY FAULT → BATTERY FAULT →  
GENERAL → REPEATER FAULT → FAULT → FAULT →

START/STOP  
SOUNDERS

SILENCE  
BUZZER

RESET

LED TEST

OVERSAMPLING  
DELAY

#### CONTROLS



#### ZONAL INDICATIONS

Zone 1 → Zone 2 → Zone 3 → Zone 4 → Zone 5 → Zone 6 → Zone 7 → Zone 8 → Zone 9 → Zone 10 →  
Zone 11 → Zone 12 → Zone 13 → Zone 14 → Zone 15 → Zone 16 → Zone 17 → Zone 18 → Zone 19 → Zone 20 →  
Zone 21 → Zone 22 → Zone 23 → Zone 24 → Zone 25 → Zone 26 → Zone 27 → Zone 28 → Zone 29 → Zone 30 →

#### ENGINEER FUNCTIONS

PRESS TO SELECT  
MODE OF TEST

FUNCTION

PRESS TO SELECT  
MODE OF SILENCE

SCROLL

PRESS TO  
CONFIRM ACTION

SELECT

GENERAL TEST

GENERAL DISABLING

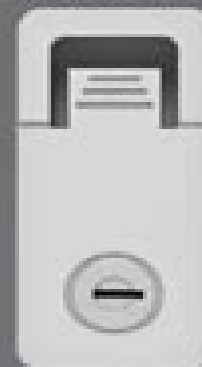
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**N Colangelo**



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calculate the number of moles from each of the following a  $6.95 \times 10^{23}$  atoms of cu b  $7.12 \times 10^{21}$  molecules of co c  $3.33 \times 10^{22}$  ions of pb d  $7.00 \times 10^{23}$  formula units of lino

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the quantity of heat for a process is represented by the letter q the sign of q for an endothermic process is positive because the system is gaining heat a chemical reaction or physical

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one milligram is equal to 0 001 grams 790 drag each number to the correct location sort the numbers based on the number of significant figures they have one significant figure 1 000

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the quantities commonly encountered in chemical problems include the number of moles of a substance the number of atoms molecules or formula units of a substance and the mass in

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it is always a good idea to estimate the answer before you do the actual calculation in this case the mass given 35 00 g is less than the molar mass so the answer should be less than 1

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jul 18 2022 chemical equations also provide us with the relative number of particles and moles that react to form products in this section you will explore the quantitative relationships that

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volume amount number of by counting a the

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the atomic mass of an element expressed in grams is the mass of a mole of the element the mass of a mole of an element is the molar mass to calculate the molar mass of a compound

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the amount of a substance that contains  $6.02 \times 10^{23}$  representative particles of that substance the amount of a substance that contains  $6.02 \times 10^{23}$  representative particles of that

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