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# **PURE MATHEMATICS SYLLABUS**

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**2015 - 2022**

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# Zimsec Pure Mathematics A 9164 2009

**Albert A Gayle**



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buffers typically consist of an acid base pair with the acid and base differing by the presence or absence of a proton a conjugate acid base pair for instance one of the buffers that maintain the ph of human blood involves carbonic acid  $\text{H}_2\text{CO}_3$



and its conjugate base the bicarbonate ion  $\text{HCO}_3^-$

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an acid and a base react not to produce a salt and a solvent but to form a new acid and a new base the concept of neutralization is thus absent brønsted lowry acid base behavior is formally independent of any solvent making it more all encompassing than the arrhenius model

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their own unique chemistries the acid and base cancel each other's chemistry to produce a rather innocuous substance water in fact the general reaction between an acid and a base is  $\text{acid} + \text{base} \rightarrow \text{water} + \text{salt}$

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