

Q2.
(2ai)

I: Deliquescence: These are compounds that tend to absorb so much water from the atmosphere that they dissolve in it to form an aqueous solution of the compounds.

II: Efflorescence: These are compounds that when exposed to the atmosphere, give off part or all of their water of crystallisation spontaneously to form a lower hydrated or anhydrous salt.



(2aii) Epick only 1 each J

- * CaCl_2 , NaOH , KOH , MgCl_2 Epick one
- * $\text{Na}_2\text{CO}_3 \cdot 10\text{H}_2\text{O}$, $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$ Epick one

(2bi)

ionization energy: This is the energy required to remove a valence electron from an atom of the element to form an ion.

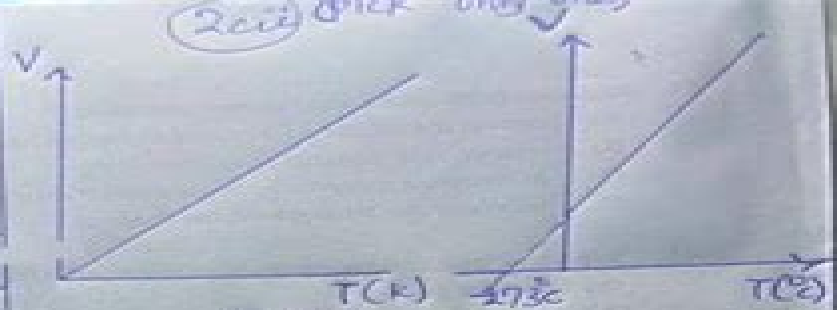
(2bii)

The second ionization energy of sodium is greater than the first ionization energy because the second electron is being removed from a positively charged sodium ion (Na^+) rather than from the neutral sodium atom. The removal of the second electron requires even more energy as it removes an electron from a stable lower energy level.

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(2ci)

Charles Law states that the volume of a fixed mass of gas varies directly with the absolute temperature provided the pressure remains constant.



(2cii)

$$V_1 = 150 \text{ cm}^3 \quad V_2 = ? \quad T_1 = 21 + 273 = 294 \text{ K}$$

$$T_2 = 45 + 273 = 318 \text{ K}$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad ; \quad \frac{150}{294} = \frac{V_2}{318} \quad ; \quad 294V_2 = 150 \times 318$$

$$\frac{294V_2}{294} = \frac{47700}{294} \quad ; \quad V_2 = 162.24 \text{ cm}^3$$

(2di)

(i) Sodium oxide



(ii) phosphorus(III) oxide:



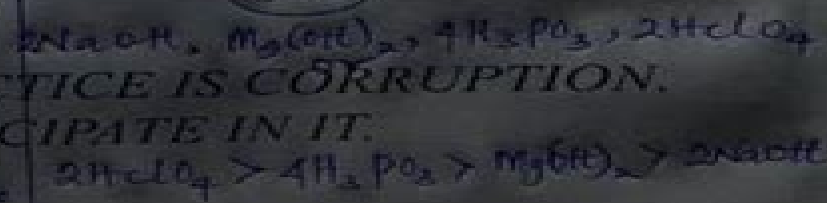
(iii) chlorine(VII) oxide:



(iv) Magnesium oxide



(2dii)



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Y Pai



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