

Enthalpy change for the thermal decomposition of potassium hydrogen carbonate =  
 $5.38 \text{ KJ/mol} + 0.32 \text{ KJ/mol} = 5.7 \text{ KJ/mol}$

**Enthalpy change for the thermal decomposition of potassium hydrogen carbonate = 5.7 KJ/mol**

4. When  $\text{KHCO}_3$  is heated, it decomposes according to the following chemical reaction:



Apply Hess' Law and use the results from questions 2b and 3b to calculate the enthalpy change for the thermal decomposition of potassium hydrogen carbonate.

# Standard Enthalpy Change Of Potassium Hydrogen Carbonate

**D Kirk**

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape with a gradient, partially overlapping the bar's end.

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**Revise A2 Chemistry for AQA** Paddy Gannon, 2005-04 Part of the series of AS and A2 revision guides this title gives students what they need to know for the AQA exams It includes material organised into bite-sized chunks of information

The National Standard Dispensatory ... Hobart Amory Hare, 1905

*A Textbook of Physical Chemistry, 6th Edition* Sharma K.K. & Sharma L.K., 2016 A Textbook of Physical Chemistry

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*Merck's Report* Theodore Weicker, 1902

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