Collins CSEC® Chemistry Workbook answers

A1 States of matter

- L 4) I) Ammonium chloride (II)
 - 10 Diffusion

Diffusion is the movement of particles from an arts of higher concentration to an arts of lower concentration until the particles are evenly distributed.

iii) The ammonia solution gave off ammonia gas and the hydrochloric acid gave off hydrogen chloride gas. The particles of the two gases diffused along the tube, collided and reacted to form ammonium chloride.







- iii) The distilled water had a higher water content than the cytoplasm inside the paw paw cells and the cell membranes were differentially permeable. The water molecules were able to move through the membranes into the cells causing the cells to swell and the strips to increase in volume. (3)
- iii) Osmosis (I)
- c) Sodium chloride draws water out of the cells of the fish and any micro-organisms by comocis. Water is, therefore, not available for chemical reactions in the cells, some of which cause decay and micro-organisms can't grow. (2)

25 40

| Property | Solid | Liquid | Gas |
|-----------------------------|---|---|---|
| Volume | | definite | variable, the volume is the same as the entire container |
| Arrangenent of particles | packed closely together, usually in a regular way | | randomly arranged with large spaces between |
| Leergy of particles | have very small amounts of kinesic energy | here medium amounts of kinetic energy | |

- b) The particles have large spaces between them, so they can be easily pushed closer together. (1)
 - The particles are packed tightly together with very little empty space between.
 - (iii) The particles move around rapidly and have weak. Forces of attraction between them, so they spread out to 6 II any available space in the container. (2)

| (1) | 40 | A: Melting | By Boiling or evaporation | |
|-----|------|---------------|------------------------------|----------|
| | | C) Preming | Di Condensation | 0.40 |
| | 64.9 | Soding or our | bion dioxide or ammonium chi | lor idle |
| | | or naphthale | mig' | 6.83 |
| (4) | 40 | Liquid | | 0.00 |
| | 100 | 54 °C | | 0.00 |
| | 8880 | COD | | (3.0) |
| | | CYYY) | | |
| | | CYY | | |

A2 Mixtures and separations

L 40 0

| | Pure substance | -04 inchains |
|-------------|--------------------|---|
| Composition | | wariable |
| Properties | Fixed and constant | variable, the components retain their individual properties |

- W) An element is a pure substance that cannot be broken down into any simpler substances by any ordinary chemical or physical means. A compound is a pure substance that contains two or more different types of element that are bonded together chemically in fixed proportions and in such a way that their properties have changed. (2)
- b) ii) The particles in a suspension are larger than those in a colloid. (I
 - The particles in a suspension settle if left undisturbed, whereas the particles in a colloid never settle.
 - iii) Light does not pass through a suspension, whereas most colloids scatter light. (1)
- A solution is a homogeneous mixture consisting of two or more components, one of which is usually a liquid.
 - Solobility is the mass of solote that will saturate 100 g of solvent at a given temperature. (1)
- d) i) 17 gper 100 gwater (1)
- 8) 45 °C (1)
 - HI) At 76 °C, 54 g of Z saturate 100 g water At 10 °C, 12.5 g of Z saturate 100 g water .: mass of Z crystallising out of a saturated solution containing 100 g water = 54 - 12.5 g = 41.5 g
- W) At 62 °C, 39 g of Z saturate 100 g water
 - ... at 62 °C, 45 g of Z saturate 100 × 45 g water = 115.4 g of water (2)
- v) At 55 °C, 33 g of Z saturate 100 g water
 - .: at 55 °C, ³³/₁₀₀ × 350 g of Z tattarate 350 g water = 115.5 g of Z. (2)

0.30

States Of Matter Chemistry Workbook Answers

Jin-Ying Zhang

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