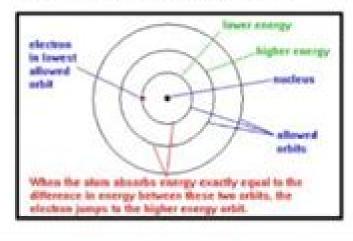
The Bohr Model

A one-dimensional model, one quantum number (n) to describe the distribution of electrons in the atom.





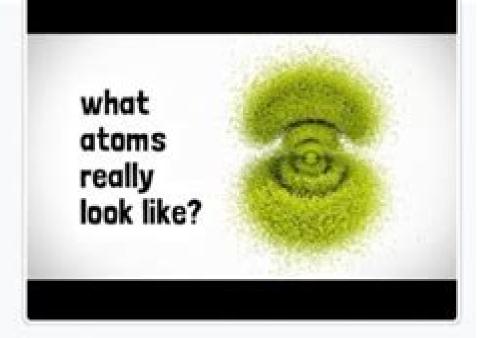




Angular (C): describes the shape of the orbital.

- Orbitals have shapes that are best described as spherical (t = 0), polar (t = 1), or cloverleaf (t = 2).
- They can even take on more complex shapes as the value of the angular quantum number becomes larger.

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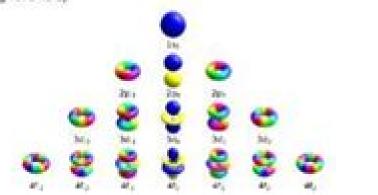
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Quantum electron configuration

Magnetic (m): describes the orientation in space

- There is only one way in which a sphere (f = 0) can be oriented in space.
- (it is called the magnetic quantum number because the effect of different orientations of orbitals was first observed in the presence of a magnetic field)



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