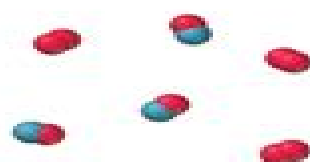


# Chapter 12 Question

Question 3: The following diagrams represent different initial concentrations of reactants for the reaction shown. Which mixture will have the fastest initial rate?



- A. A
- B. B
- C. C
- D. They all have the same initial rate.



(a)



(b)



(c)

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Question 6: What is the reaction order for peroxydisulfate ion ( $\text{S}_2\text{O}_8^{2-}$ )?



- A. 0
- B. 1
- C. 2
- D. 3

Exp't	$[\text{S}_2\text{O}_8^{2-}]$ , M	$[\text{I}^-]$ , M	Initial Rate M/s
1	0.080	0.034	$2.2 \times 10^{-4}$
2	0.080	0.017	$1.1 \times 10^{-4}$
3	0.16	0.017	$2.2 \times 10^{-4}$

Question7: What is the reaction order for iodide ion ( $\text{I}^-$ )?

# Teaching Transparency Worksheet Phase Diagrams Answers

**G Orfield**



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