

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 1) Consider the following reaction:



The average rate of appearance of B is given by $\Delta[B]/\Delta t$. Comparing the rate of appearance of B and the rate of disappearance of A, we get $\Delta[B]/\Delta t = \underline{\hspace{1cm}} \times (-\Delta[A]/\Delta t)$.

- A) $-2/3$ B) $-3/2$ C) $+2/3$ D) $+3/2$ E) $+1$
- 2) A burning splint will burn more vigorously in pure oxygen than in air because
- A) nitrogen is a reactant in combustion and its low concentration in pure oxygen catalyzes the combustion.
 B) oxygen is a reactant in combustion and the concentration of oxygen is higher in pure oxygen than it is in air.
 C) oxygen is a product of combustion.
 D) nitrogen is a product of combustion and the system reaches equilibrium at a lower temperature.
 E) oxygen is a catalyst for combustion.
- 3) Of the following, all are valid units for a reaction rate except _____.
- A) mol/L B) M/s C) mol/hr D) mol/L-hr E) g/s

A flask is charged with 0.124 mol of A and allowed to react to form B according to the reaction $A(g) \rightarrow B(g)$. The following data are obtained for [A] as the reaction proceeds:

Time (s)	1	10	20	30	40
Moles of A	0.124	0.110	0.088	0.073	0.054

- 4) The average rate of disappearance of A between 10 s and 20 s is _____ mol/s.
- A) 1.1×10^{-3} B) 4.4×10^{-3} C) 454 D) 9.90×10^{-3} E) 2.2×10^{-3}

The peroxydisulfate ion ($S_2O_8^{2-}$) reacts with the iodide ion in aqueous solution via the reaction:



An aqueous solution containing 0.050 M of $S_2O_8^{2-}$ ion and 0.072 M of I^- is prepared, and the progress of the reaction followed by measuring $[I^-]$. The data obtained is given in the table below.

Time (s)	0	400	800	1200	1600
$[I^-]$ (M)	0.072	0.057	0.046	0.037	0.029

- 5) The concentration of $S_2O_8^{2-}$ remaining at 800 s is _____ M.
- A) 0.015 B) 0.041 C) 0.076 D) 4.00×10^{-3} E) 0.046
- 6) Of the units below, _____ are appropriate for a first-order reaction rate constant.
- A) $L \cdot mol^{-1} s^{-1}$ B) s^{-1} C) $M^{-1} s^{-1}$ D) mol/L E) $M s^{-1}$

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