

# **JOHN DEERE**

## **106 TECHNICAL MANUAL**

### **TM1206**

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# Tm1206 Technical Manual

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 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$   $\Delta h = 71.2 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{g})$   $\Delta h = 283 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g})$

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 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ}$  mole  
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ}$  mole  
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ}$  mole

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reversed the sign of the  $\Delta h$  value is also reversed two if an equation is multiplied by a

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{O}$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta h$  55.8 kJ as shown note that the enthalpy change  $\Delta h_{\text{products}} - \Delta h_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta h_{\text{products}} > \Delta h_{\text{reactants}}$

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