UNEB UACE CHEMISTRY PAPER 2 2018

SECTION A

 a) A compound Q contains 64.9% carbon, 13.5% hydrogen and the rest being oxygen.
85g of Q in the vapour. form occupied 969.8cm3 at 200°C.

i) Calculate the empirical formula of Q ii) Determine the molecular formula of Q

(The molar gas constant, $R \approx 8.31JK^{'I}$.mol $^{'I}$)

b) Q reacts with sodium with offervescence but has no effect on sodium carbonate. Write the names and the structures of all possible isomers of Q

c) When treated with anhydrous zinc chloride in the presence of concentrated hydrochloric acid. Q formed two layers after about 8 minutes. Identity Q.

d) Q reacted with acidified chromium trioxide to give a compound A. Write equation for the reaction:

i) leading to the formation of A.

ii) between A and acidified 2, 4-dinitrophenylbdrazine and outline a mechanism for the reaction.

e)Write equation to shoe how Q can be prepared from an alkene and outline a mechanism for the reaction.

2. a) Explain the principle on which stram distillation is based

bb State

the conditions necessary for steam distillation.

ii) one advantage of steam distillation over fractional distillation.

c) The following data was obtained for the steam distillation of bromobenzene at 760mmHg pressure

Temperature ("C)	90	94	94	96	98	TOO
Vapour pressure of water (mmHg)	5.26	100	64.1	6558	707	760:
Vapor pressure of bromobenzene (mmHg)	17.	106	114	123	132	141

i) On the same axes, plot graphs of vapour possure of the mixture and of each component against temperature.

ii) Determine the boiling point of the mixture

iii) Calculate the percentage by man of bromobenze in the distillate.

d) State how a sample of dry bromobening can be obtained from the distillate

e) State one other application steam distillation.

3.a) Explain what is meant by the term.

i) Solubility product

ii) Common ion effect.

b)ii) describe briefly bow the solubility product of magnesium hydroxide in water can be determine.

ii) State hoe solubility product can be used to predict precipitation or dissolution.

c) The solubility product of calicium sulphate at 25°C is 2.4 × 10° mol² dm 6. Calculate the solubility of calcium sulphate in mol dm 3 in:

D water

ii) 0.5M sulphuric acid

d) State one condition under which solubility product is invlaid.

The table below shows the melting point of some exides of elements in Group (IV) of the Periodic Table.

Compound	CO ₂	SiO ₂	GeO ₂	SmO ₂	PbO ₂	
Melting point -56.5 (°C)		1700	116	1827	752	

a) Explain the trend in the melting points of the oxides

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