

Titration Practice Worksheet

Find the requested quantities in the following problems:

- 1) If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl?

$$\begin{aligned} M_1 &= 0.1 \text{ M} \\ V_1 &= 54 \text{ mL} \\ M_2 &= ? \\ V_2 &= 125 \text{ mL} \end{aligned}$$

$$\begin{aligned} M_1 V_1 &= M_2 V_2 \\ (0.1 \text{ M})(54 \text{ mL}) &= (125 \text{ mL}) M_2 \\ M_2 &= \frac{(0.1 \text{ M})(54 \text{ mL})}{125 \text{ mL}} = 0.0432 \text{ M} = 4.32 \times 10^{-2} \text{ M} \end{aligned}$$

The concentration of HCl is $4.32 \times 10^{-2} \text{ M}$.

- 2) If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?

$$\begin{aligned} M_1 &= 0.05 \text{ M} \\ V_1 &= 25 \text{ mL} \\ M_2 &= ? \\ V_2 &= 345 \text{ mL} \end{aligned}$$

$$\begin{aligned} M_1 V_1 &= M_2 V_2 \\ (0.05 \text{ M})(25 \text{ mL}) &= (345 \text{ mL}) M_2 \\ M_2 &= \frac{(0.05 \text{ M})(25 \text{ mL})}{345 \text{ mL}} = 0.00362 \text{ M} = 3.62 \times 10^{-3} \text{ M} \end{aligned}$$

The concentration of NaOH is $3.62 \times 10^{-3} \text{ M}$.

- 3) If it takes 50 mL of 0.5 M KOH solution to completely neutralize 125 mL of sulfuric acid solution (H_2SO_4), what is the concentration of the H_2SO_4 solution?



$$n = MV$$

$$n = (0.5 \text{ M})(50 \text{ mL}) = 25 \text{ mol OH}^-$$

2 times $[\text{OH}^-]$ is needed

to neutralize $[\text{H}^+]$, $\rightarrow [\text{H}^+] = \frac{1}{2} [\text{OH}^-]$

$$M = \frac{12.5 \text{ mol}}{125 \text{ mL}} = \boxed{0.10 \text{ M}}$$

- 4) Can I titrate a solution of unknown concentration and still get a meaningful answer? Explain your answer in a few sentences.

Titration cannot be done without molarity of at least one of the substances. In order to solve $M_1 V_1 = M_2 V_2$, both molarities cannot be unknown.

- 5) Explain the difference between an endpoint and equivalence point in a titration.

Endpoint: titration can be stopped.

Equivalence point: when $[\text{H}^+]$ and $[\text{OH}^-]$ are neutralized.

The effectiveness of a titration is measured by the closeness of the endpoint to the equivalence point.

Titration Nursing Math Problems With Key

Michael Seilmaier



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pode propiciar para o bem e para o mal das crianças dessa idade. O escolar apresenta o desenvolvimento típico da criança em idade escolar, os mecanismos para tornar esse processo saudável, os sintomas de situações de alerta e os principais fatores de risco observados ao desenvolvimento saudável. Reafirmamos que o conteúdo desta obra também é resultante da integração dos efeitos promovidos pelos ambientes social, biológico e psicológico, demonstrando a importância de abordagem integral. Exemplificamos com o clássico da literatura e abordamos novos problemas da modernidade cuja literatura esteja crescente, como os tempos de exposição a telas e o uso de jogos nessa faixa etária. Este livro condensa evidências e experiências de autores engajados no desenvolvimento humano e na otimização de potenciais de capital mental. Parte da coleção Desenvolvimento Humano no Ciclo da Vida, esta obra busca apresentar as etapas da vida com suas particularidades, mas com a perspectiva contnua de que todas as fases devem ser compreendidas e trabalhadas de forma a otimizar os potenciais para o indivíduo e para a sociedade. Profissionais experientes tecem comentários sobre problemas atuais e informam dentro do melhor que a literatura científica sugere como abordar a criança dessa faixa etária de forma a evitar déficits e aprimorar a trajetória de seu desenvolvimento, otimizando o capital mental do indivíduo e da sociedade.

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