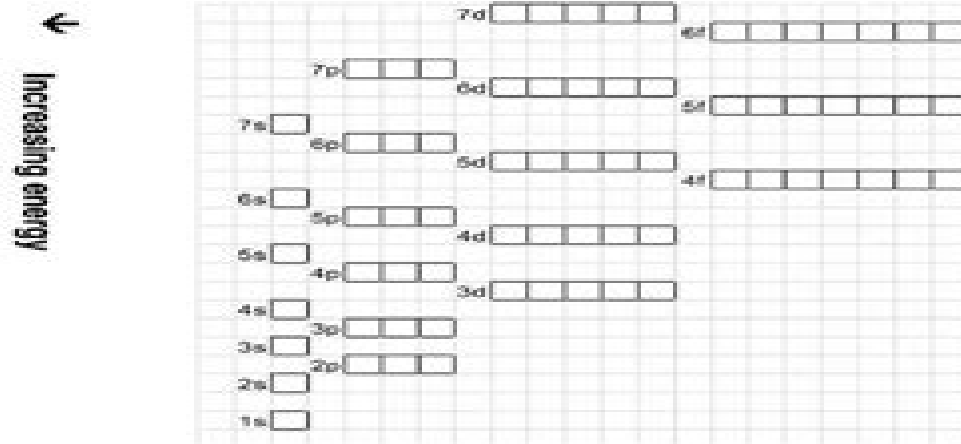


APARTMENT ELECTRON CONFIGURATION

Background Information: The arrangement of electrons within the orbitals of an atom is known as the electron configuration. The most stable arrangement is called the ground-state electron configuration. This is the configuration where all of the electrons in an atom reside in the lowest energy orbitals possible. Keeping in mind several "rules" that the electrons must obey, we are able to predict the electron configurations of elements using the electron orbital diagrams and periodic table.

Electron Orbital Diagram



subshell can hold 2 electrons; the p can hold 6; the d can hold 10; and the f can hold 14.

Nucleus

y

The distributions of orbitals can be laid out in the electron orbital diagram pictured to the left. You read it from the bottom (or closest to the nucleus) to the top in order of increasing energy.

The bottom energy level is level 1 - it has the lowest energy. Each  represents an orbital. You can see that there is 1 orbital for an s subshell. There are 3 orbitals for a p subshell, 5 for a d, and 7 for an f subshell. **Each orbital can hold 2 electrons.** Therefore, the s

Subshell	Number of Orbitals	Shape of orbital	Max number of electrons
s	1	Sphere	$1 \times 2 = 2$
p	3	Dumbbells	$3 \times 2 = 6$
d	5	Clover	$5 \times 2 = 10$

Understanding Electron Configuration Lab Answer Key

KJ Lindholm-Leary



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