CHEMISTRY FORM 4 SCHEMES OF WORK TERM 1,2,3

	SCHEME OF WORK			FORM FOUR CHEMISTRY	TERM ONE		
9K 8O	L/ NO	TOPIC / SUBTOPIC	LESSON / SPECIFIC OBJECTIVES	TEACHING / LEARNING ACTIVITIES	MATERIALS / RESOURCES	REFERENCES	REMARKS
	1,	ACIDS, BASES AND SALTS. Strength of acids. Acids in equeous form.	By the end of the leason, the learner should be oble for Define an acid in terms of hydrogen ions. Explain strength of acids in equeous form in terms of number of hydrogen ions present.	Class experiments: investigate reactions of magnesium and sinc carbonate with different scids. Malco and record observations in tabular form. Malco deductions from the observations and ionic equations. Detailed discussion leading to the definition of an acid and explanation of strength of an acid.	Magnesium strip, pinc carbonate, 2M HCL 2M H ₂ SO ₄ , 2M ethanoic acid.	KLB BK IV Pages 1-4	
	3,	p.H values of acids. Electrical conductivities of aqueous acids.	Determine strength of soids using pH values. Determine strengths of soids by comparing their electrical conductivities. Classify soids as either strong or weak in terms of partial dissociations in squeous solutions.	Q/A: seview determination of strength of acids using a literate paper and pill scale. Class / group experiments: second colour of universal indicator in 2M HCl and 2M ethanoic scal. Set up with smothers of 2M HCl and 2M ethanoic noid in turns. Record associates of current. Discuss the observations. Write corresponding some equations.	Universal indicator, 2M HCl, 2M ethanoic acid, dry cells, carbon electrodes, milli-ammeters, wires, switches etc.	KLB BK IV Pages 4-6	
	5	Definition of a base in terms of hydroxide ions.	Define a base in terms of hydroxide ions.	Teacher demonstration: Dissofre-calcium hydronide in water. Carry out literal test on the resulting solution. Discuss the results; hence define a hase in terms of hydronide tons.	Red litmus peper, calcium hydroxide solid.	KLB BK IV Pages 6-7	

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LP Steffe

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