STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

$$c = f\lambda$$

$$E_{photon} = hf$$

$$E_{proton} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

$$P_{\rm T} = P_1 + P_2 + P_3 + \dots$$

$$PV = nRT$$

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

$$P_1V_1 = P_2V_2$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

$$M = \frac{\text{mol}}{1}$$

$$K_{w} = [H^{*}][OH^{-}]$$

$$Volume of (molarity of solution 1) (molarity of solution 2) (molarity of solution 2)$$

$$V_1 M_1 = V_2 M_2$$

$$pH = -log[H^*]$$

THERMOCHEMISTRY

$$Q = mc_p \Delta T$$

Enthalpy of
$$=$$
 $\begin{pmatrix} enthalpy \\ of products \end{pmatrix} - \begin{pmatrix} enthalpy \\ of reactants \end{pmatrix}$

$$\Delta H = \Delta H_f^0(\text{products}) - \Delta H_f^0(\text{reactants})$$

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