

STAAR CHEMISTRY REFERENCE MATERIALS



ATOMIC STRUCTURE

Speed of light = (frequency)(wavelength)

$$c = f\lambda$$

Energy = (Planck's constant)(frequency)

$$E_{\text{photon}} = hf$$

Energy = $\frac{(\text{Planck's constant})(\text{speed of light})}{(\text{wavelength})}$

$$E_{\text{photon}} = \frac{hc}{\lambda}$$

BEHAVIOR OF GASES

Total pressure of a gas = $\left(\begin{array}{l} \text{sum of the partial pressures} \\ \text{of the component gases} \end{array} \right)$

$$P_T = P_1 + P_2 + P_3 + \dots$$

(Pressure)(volume) = (moles)(ideal gas constant)(temperature)

$$PV = nRT$$

$\frac{(\text{Initial pressure})(\text{initial volume})}{(\text{Initial moles})(\text{initial temperature})} = \frac{(\text{final pressure})(\text{final volume})}{(\text{final moles})(\text{final temperature})}$

$$\frac{P_1 V_1}{n_1 T_1} = \frac{P_2 V_2}{n_2 T_2}$$

(Initial pressure)(initial volume) = (final pressure)(final volume)

$$P_1 V_1 = P_2 V_2$$

$\frac{(\text{Initial volume})}{(\text{Initial temperature})} = \frac{(\text{final volume})}{(\text{final temperature})}$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

$\frac{(\text{Initial volume})}{(\text{Initial moles})} = \frac{(\text{final volume})}{(\text{final moles})}$

$$\frac{V_1}{n_1} = \frac{V_2}{n_2}$$

SOLUTIONS

Molarity = $\frac{\text{moles of solute}}{\text{liter of solution}}$

$$M = \frac{\text{mol}}{\text{L}}$$

Ionization constant of water = $\left(\begin{array}{l} \text{hydrogen ion} \\ \text{concentration} \end{array} \right) \left(\begin{array}{l} \text{hydroxide ion} \\ \text{concentration} \end{array} \right)$

$$K_w = [\text{H}^+][\text{OH}^-]$$

$\left(\begin{array}{l} \text{Volume of} \\ \text{solution 1} \end{array} \right) \left(\begin{array}{l} \text{molarity of} \\ \text{solution 1} \end{array} \right) = \left(\begin{array}{l} \text{volume of} \\ \text{solution 2} \end{array} \right) \left(\begin{array}{l} \text{molarity of} \\ \text{solution 2} \end{array} \right)$

$$V_1 M_1 = V_2 M_2$$

pH = -logarithm (hydrogen ion concentration)

$$\text{pH} = -\log[\text{H}^+]$$

THERMOCHEMISTRY

Heat gained or lost = (mass) $\left(\begin{array}{l} \text{specific} \\ \text{heat} \end{array} \right) \left(\begin{array}{l} \text{change in} \\ \text{temperature} \end{array} \right)$

$$Q = mc_p \Delta T$$

Enthalpy of reaction = $\left(\begin{array}{l} \text{enthalpy} \\ \text{of products} \end{array} \right) - \left(\begin{array}{l} \text{enthalpy} \\ \text{of reactants} \end{array} \right)$

$$\Delta H = \Delta H_f^\circ(\text{products}) - \Delta H_f^\circ(\text{reactants})$$

Staar Chemistry Reference Materials Scavenger Hunt

Responsive Education Solutions

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