



Student Workbook

Exambusters PRAXIS 2 Chemistry

Basic Review Flash Cards

ACIDS & BASES...605 of 720

In a saturated solution of BaSO_4 ,
 $[\text{Ba}^{2+}] = 2 \times 10^{-5}$,
find K_{sp} .

**Sample
Flash Card?**

Answer 605 of 720

Each molecule of BaSO_4 that ionizes produces equal concentrations of ions; therefore: $[\text{Ba}^{+2}] = [\text{SO}_4^{-2}]$

$$K_{sp} = [\text{Ba}^{+2}][\text{SO}_4^{-2}] = [2 \times 10^{-5}][2 \times 10^{-5}] = 4 \times 10^{-10}$$

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