CHAPTER (1) STUDY GUIDE FOR CONTENT MASTERY

#### Section 12.3 Limiting Reactants

In your treatment, reself about only reactions stop and flow to determine the limiting reactions

Study the diagram develop a chemical reaction and the chemical equation that reprement the reaction. These complete the table. Now your calculations for questions 25–27 in the space below the table.

The mode momen of G<sub>p</sub> NGs and NG<sub>p</sub> are 12.00 plant. NGS plant, and 40.01 plant, or provided as

Annual of S <sub>p</sub>	Assured of the	Association (China	Sintley Market	Amount and Komm of Season Resistant
Employee	- I melles des	- I moltoubo	1000	recta:
Emilional Control	Armetecphre	A rectinguist	40	3 miles de O <sub>2</sub>
Condition	E-incliniques	1. A molecules	2.0,	3. 4 melecules NO
1.00 met	2-10 mad	4. 2.00 mail	S. none	6. none
4.00 (4.0	0.00 mol	2, 4.00 mol	A. NO	s. 2.60 mol O <sub>2</sub>
5.00 mad	100 med	16. 7.00 mol.	11. 500	12. 1.50 mail O <sub>2</sub>
1.00 med	140 mt	13. 2.00 mot	sa. O <sub>3</sub>	16. 2.00 mol NO
1.000 met	4,000 med	16. 0.200 mol	17. 160	18. 0.400 mol O <sub>3</sub>
18.00 g	60.00	16. 92.02 g	29. none	25. none
16.00 g	min.	33. 65.01 g	20. 03	34. 50.12 g NO
1000	300 May 1	26. 28.74 g	ps. O <sub>1</sub>	29. 1.34 g 190

belienced equation male ratio = 2 mol MO/1 mol  $\Omega_2$ 

10.00 g-0; × 1 mol Oy/12.00 g-0; + 0.3125 mol O<sub>3</sub>

20.00 p.NO × 1 mol NO/NO.01 p.NO = 0.6664 mol NO

actual mole ratio = 0.6664 mol NO/0.3125 mol  $O_{\chi}$  = 2.132 mol NO/1.000 mol  $O_{\chi}$  Because the actual mole ratio of NO· $O_{\chi}$  is larger than the balanced equation mole ratio of NO· $O_{\chi}$  there is an excess of NO;  $O_{\chi}$  is the limiting reactuals.

Mann of NO used = 0.3125 past 65; × 2 mol NO/1 past 65; = 0.6250 mol NO 0.6210 past NO = 30.01 p NO/1 past NO = 18.76 p NO

Mans of NO<sub>2</sub> produced = 6.6250 pask NO<sub>2</sub> = 46.01 g NO<sub>2</sub>/1 pask NO<sub>2</sub> = 28.76 g NO<sub>2</sub>. Extens NO = 20.00 g NO = 18.76 g NO = 1.24 g NO

Shade Goods for Contact Michael

Chamberly Metter and Dungs Y Drauter 16 .

1

OHAPTIR (D) STUDY GUIDE FOR CONTENT MASTERY

#### Section 52.4 Percent Yield

In your bridging, read about the plots of products

Study the diagram and the reample problem.

man of product from imperimental measurement

main of product producted from standardwares concration using

- a. Heat of traction
- b. A map many to many commences

of more man.

- 1. Write the balanced channel countries.
- Consists the number of moins of reactions, using molar main.
- Calculate the number of moles of product, using the appropriate more radio.
- the appropriate more ratio.

  4. Cultiviate the mass of product, using the reciprocal.

Example Posture: The following character apparent to production of pullines code, a solution; until in the manufacturing of some acasisoshear devices.

In one exponence, the exercise, yielded 7.47 g of the exists three a 7.95 g sample of pullson. Describes the parametrisks of this exercise. The tentar master of Co and Co<sub>2</sub>O<sub>2</sub> are 89.72 gived and 167.44 gives, respectively.

Use the information in the diagram and example problem to evaluate each value or expension below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- L arted jobt volume 7.42 q GayOu
- At these of brackets Thinly the Contract
- 8. results of make of tructure  $7.00 \pm G_0 \times \frac{46.77 \pm G_0}{1 \text{ mod } G_0}$  7.00 g Ga  $\times$  1 mod Ga/90.73 g Ga
- 4. samples of matter of product 0.00 and Ga  $\times \frac{2 \operatorname{mod} \operatorname{Ga}_{1}(\lambda)}{2 \operatorname{mod} \operatorname{Ga}} \stackrel{Q. NSS}{=} \frac{\operatorname{mod} \operatorname{Ga}_{2}(\lambda)}{2 \operatorname{mod} \operatorname{Ga}_{2}(\lambda)/4 \operatorname{mod} \operatorname{Ga}}$
- **b.** the second yield 0.0500 and  $Ca_{ij}U_{ij} \times \frac{(iG) 44 \text{ gr} Ga_{ij}U_{ij}}{1 \text{ and } Ga_{ij}U_{ij}}$  exerces.
- 4. provide profet:  $\frac{4.17 \pm G_{12}G_{1}}{2.47 \pm G_{12}G_{2}} \approx 100$   $\frac{3.42 \pm G_{12}G_{1}/8.17 \pm G_{12}G_{2} \times 100}{1.42 \pm G_{12}G_{2}/8.17 \pm G_{12}G_{2} \times 100}$
- Dentitry Matter and Dungs + Dunter 16

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