

AP Chemistry Chapter 9 Answers – Zumdahl

9.15

H_2O has $2(1) + 6 = 8$ valance electrons

H_2O has a tetrahedral arrangement of the electron pairs about the O atom that requires sp^3 hybridization. Two of the four sp^3 hybrid orbitals are used to form bonds to the two hydrogen atoms and the other two sp^3 hybrid orbitals hold the two lone pairs of oxygen. The two O-H bonds are formed from overlap of the sp^3 hybrid orbitals on oxygen with the 1s atomic orbitals on the hydrogen atoms.

9.17

H_2CO has $2(1) + 4 + 6 = 12$ valence electrons

The central carbon atom has a trigonal planar arrangement of the electron pairs which requires sp^2 hybridization. The two C-H sigma bonds are formed from overlap of the sp^2 hybrid orbitals on carbon with the hydrogen 1s atomic orbitals. The double bond between carbon and oxygen consists of one σ and one π bond. The oxygen atom, like the carbon atom, also has a trigonal planar arrangement of the electrons which requires sp^2 hybridization. The σ bond in the double bond is formed from overlap of a carbon sp^2 hybrid orbital with an oxygen sp^2 hybrid orbital. The π bond in the double bond is formed from overlap of the unhybridized p atomic orbitals. Carbon and oxygen each have one unhybridized p atomic orbital which are parallel to each other. When two parallel p atomic orbitals overlap, a π bond results.

9.19

Ethane, C_2H_6 , has $2(4) + 6(1) = 14$ valence electrons.

The carbon atoms are sp^3 hybridized. The six C-H sigma bonds are formed from overlap of the sp^3 hybrid orbitals on C with the 1s atomic orbitals from the hydrogen atoms. The carbon-carbon sigma bond is formed from overlap of an sp^3 hybrid orbital on each C atom.

Ethanol, $\text{C}_2\text{H}_5\text{O}$ has $2(4) + 6(1) + 6 = 20$ e⁻

The two C atoms and the O atom are sp^3 hybridized. All bonds are formed from overlap with these sp^3 hybrid orbitals. The C-H and O-H sigma bonds are formed from overlap of sp^3 hybrid orbitals with hydrogen 1s atomic orbitals. The C-C and C-O sigma bonds are formed from overlap of the sp^3 hybrid orbitals on each atom.

9.21

See Exercises 8.67 and 8.73 for the Lewis structures. To predict the hybridization, first determine the arrangement of electron pairs about each central atom using the VSEPR model; then utilize the information in Figure 9.24 of the text to deduce the hybridization required for that arrangement of electron pairs.

Study Guide Zumdahl Answers Ap Chemistry

Iris Stovall, Donald J. DeCoste



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Table of Contents Study Guide Zumdahl Answers Ap Chemistry

1. Understanding the eBook Study Guide Zumdahl Answers Ap Chemistry
 - The Rise of Digital Reading Study Guide Zumdahl Answers Ap Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide Zumdahl Answers Ap Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide Zumdahl Answers Ap Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide Zumdahl Answers Ap Chemistry
 - Personalized Recommendations
 - Study Guide Zumdahl Answers Ap Chemistry User Reviews and Ratings
 - Study Guide Zumdahl Answers Ap Chemistry and Bestseller Lists
5. Accessing Study Guide Zumdahl Answers Ap Chemistry Free and Paid eBooks
 - Study Guide Zumdahl Answers Ap Chemistry Public Domain eBooks
 - Study Guide Zumdahl Answers Ap Chemistry eBook Subscription Services
 - Study Guide Zumdahl Answers Ap Chemistry Budget-Friendly Options
6. Navigating Study Guide Zumdahl Answers Ap Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Study Guide Zumdahl Answers Ap Chemistry Compatibility with Devices
 - Study Guide Zumdahl Answers Ap Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Study Guide Zumdahl Answers Ap Chemistry
 - Highlighting and Note-Taking Study Guide Zumdahl Answers Ap Chemistry
 - Interactive Elements Study Guide Zumdahl Answers Ap Chemistry

8. Staying Engaged with Study Guide Zumdahl Answers Ap Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide Zumdahl Answers Ap Chemistry
9. Balancing eBooks and Physical Books Study Guide Zumdahl Answers Ap Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide Zumdahl Answers Ap Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide Zumdahl Answers Ap Chemistry
 - Setting Reading Goals Study Guide Zumdahl Answers Ap Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide Zumdahl Answers Ap Chemistry
 - Fact-Checking eBook Content of Study Guide Zumdahl Answers Ap Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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