

Use the periodic table and your knowledge of periodic trends to answer the following questions.

Atomic Radius

- Which atom in each pair has the larger atomic radius? Circle your answer.
 a) Li or K b) Ca or Ni c) Ga or B d) O or C e) Cl or Br
 f) Be or Ba g) Si or S h) Fe or Au
- Why do atoms get smaller as you move left to right in a period? Larger as you move down a group?
- Arrange the following in order of increasing atomic radius.
 Mg, Na, P, Si and Ar _____

Ionic Radius

- Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.
- Which particle has the larger radius in each atom/ion pair?
 a) Na, Na⁺ b) S, S²⁻ c) I, I⁻ d) Al, Al³⁺ e) Cl, Cl⁻

Ionization Energy

- Which element in each pair has the larger ionization energy?
 a) Na or O b) Be or Ba c) Ar or F d) Cu or Ra e) I or Ne
 f) K or V g) Ca or Fr h) W or Se
- Arrange the following groups of elements in order of increasing ionization energy.
 a) Be, Mg, Sr _____ b) Bi, Cs, Ba _____ c) Na, Al, S _____

Electronegativity

- Which element in each pair has a higher electronegativity value?
 a) Cl, F b) C, N c) Mg, Ne d) As, Ca e) Fr, Cs

Metallic Character

- Arrange the elements from least to most metallic: Sodium, barium, cesium, magnesium, calcium

Electron Affinity

- Write the chemical equation for the addition of an electron to a chlorine atom.
- Arrange the following atoms in order of **increasing** electron affinity: Br, Sb, I, Te, Cl

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Thomas Griffiths



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Table of Contents Teaching Transparency Master Atomic And Ionic Radii Answers

1. Understanding the eBook Teaching Transparency Master Atomic And Ionic Radii Answers
 - The Rise of Digital Reading Teaching Transparency Master Atomic And Ionic Radii Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Teaching Transparency Master Atomic And Ionic Radii Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Teaching Transparency Master Atomic And Ionic Radii Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Teaching Transparency Master Atomic And Ionic Radii Answers
 - Personalized Recommendations
 - Teaching Transparency Master Atomic And Ionic Radii Answers User Reviews and Ratings

- Teaching Transparency Master Atomic And Ionic Radii Answers and Bestseller Lists
- 5. Accessing Teaching Transparency Master Atomic And Ionic Radii Answers Free and Paid eBooks
 - Teaching Transparency Master Atomic And Ionic Radii Answers Public Domain eBooks
 - Teaching Transparency Master Atomic And Ionic Radii Answers eBook Subscription Services
 - Teaching Transparency Master Atomic And Ionic Radii Answers Budget-Friendly Options
- 6. Navigating Teaching Transparency Master Atomic And Ionic Radii Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Teaching Transparency Master Atomic And Ionic Radii Answers Compatibility with Devices
 - Teaching Transparency Master Atomic And Ionic Radii Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Teaching Transparency Master Atomic And Ionic Radii Answers
 - Highlighting and Note-Taking Teaching Transparency Master Atomic And Ionic Radii Answers
 - Interactive Elements Teaching Transparency Master Atomic And Ionic Radii Answers
- 8. Staying Engaged with Teaching Transparency Master Atomic And Ionic Radii Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Teaching Transparency Master Atomic And Ionic Radii Answers
- 9. Balancing eBooks and Physical Books Teaching Transparency Master Atomic And Ionic Radii Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Teaching Transparency Master Atomic And Ionic Radii Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Teaching Transparency Master Atomic And Ionic Radii Answers
 - Setting Reading Goals Teaching Transparency Master Atomic And Ionic Radii Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Teaching Transparency Master Atomic And Ionic Radii Answers
 - Fact-Checking eBook Content of Teaching Transparency Master Atomic And Ionic Radii Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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