

Use the periodic table and your knowledge of periodic trends to answer the following questions.

**Atomic Radius**

- Which atom in each pair has the larger atomic radius? Circle your answer.  
 a) Li or K      b) Ca or Ni      c) Ga or B      d) O or C      e) Cl or Br  
 f) Be or Ba      g) Si or S      h) Fe or Au
- Why do atoms get smaller as you move left to right in a period? Larger as you move down a group?
- Arrange the following in order of increasing atomic radius.  
 Mg, Na, P, Si and Ar \_\_\_\_\_

**Ionic Radius**

- Explain the relationship between the relative size of an ion to its neutral atom and the charge on the ions.
- Which particle has the larger radius in each atom/ion pair?  
 a) Na, Na<sup>+</sup>      b) S, S<sup>2-</sup>      c) I, I<sup>-</sup>      d) Al, Al<sup>3+</sup>      e) Cl, Cl<sup>-</sup>

**Ionization Energy**

- Which element in each pair has the larger ionization energy?  
 a) Na or O      b) Be or Ba      c) Ar or F      d) Cu or Ra      e) I or Ne  
 f) K or V      g) Ca or Fr      h) W or Se
- Arrange the following groups of elements in order of increasing ionization energy.  
 a) Be, Mg, Sr \_\_\_\_\_      b) Bi, Cs, Ba \_\_\_\_\_      c) Na, Al, S \_\_\_\_\_

**Electronegativity**

- Which element in each pair has a higher electronegativity value?  
 a) Cl, F      b) C, N      c) Mg, Ne      d) As, Ca      e) Fr, Cs

**Metallic Character**

- Arrange the elements from least to most metallic: Sodium, barium, cesium, magnesium, calcium

**Electron Affinity**

- Write the chemical equation for the addition of an electron to a chlorine atom.
- Arrange the following atoms in order of **increasing** electron affinity: Br, Sb, I, Te, Cl

# Teaching Transparency Worksheet 20 Atomic And Ionic Radii

**PT Brinkman**



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