

CHAPTER 4 REVIEW*Arrangement of Electrons in Atoms***SECTION 1**

SHORT ANSWER Answer the following questions in the space provided.

1. In what way does the photoelectric effect support the particle theory of light?

In order for an electron to be ejected from a metal surface, the electron must be struck by a single photon with at least the minimum energy needed to knock the electron loose.

2. What is the difference between the ground state and the excited state of an atom?

The ground state is the lowest energy state of the atom. When the atom absorbs energy, it can move to a higher energy state, or excited state.

3. Under what circumstances can an atom emit a photon?

A photon is emitted when an atom moves from an excited state to its ground state or to a lower-energy excited state.

4. How can the energy levels of the atom be determined by measuring the light emitted from an atom?

When an atom loses energy, it falls from a higher energy state to a lower energy state. The frequency of the emitted light, observed in an element's line-emission spectrum, may be measured. The energy of each transition is calculated using the equation $E = h\nu$, where ν is the frequency of each of the lines in the element's line-emission spectrum. From the analysis of these results, the energy levels of an atom of each element may be determined.

5. Why does electromagnetic radiation in the ultraviolet region represent a larger energy transition than does radiation in the infrared region?

Energy is proportional to frequency, and ultraviolet radiation has a higher frequency than infrared radiation. To produce ultraviolet radiation, electrons must drop to lower energy levels than they do to produce infrared radiation.

The Arrangement Of Electrons In An Atoms Worksheet Answers

George Graybill



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and electron affinity and discuss how these trends are related to an element's position on the table. In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases. We will also explore the importance of understanding atomic structure in fields such as biochemistry, environmental science, and materials science. Throughout this book we will provide clear explanations, helpful illustrations, and interactive activities to reinforce key concepts. Whether you are a visual learner, a hands-on learner, or a theoretical learner, this book is designed to meet your needs. We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry. So let's dive in and unlock the mysteries of the atom together.

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