

# Physics

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(Chapter 12) (Kinetic Theory) (Exercises)  
(Class 11)

## Question 12.1:

Estimate the fraction of molecular volume to the actual volume occupied by oxygen gas at STP. Take the diameter of an oxygen molecule to be  $3 \text{ \AA}$ .

## Answer 12.1:

Diameter of an oxygen molecule,  $d = 3 \text{ \AA}$

Radius,  $r = \frac{d}{2} = \frac{3}{2} \text{ \AA} = 1.5 \text{ \AA} = 1.5 \times 10^{-8} \text{ cm}$

Actual volume occupied by 1 mole of oxygen gas at STP =  $22400 \text{ cm}^3$

Molecular volume of oxygen gas,  $V = \frac{4}{3} \pi r^3 N$

Where,  $N$  is Avogadro's number =  $6.023 \times 10^{23}$  molecules/mole

$$\therefore V = \frac{4}{3} \times 3.14 \times (1.5 \times 10^{-8})^3 \times 6.023 \times 10^{23} = 8.51 \text{ cm}^3$$

$$\text{Ratio of the molecular volume to the actual volume of oxygen} = \frac{8.51}{22400} = 3.8 \times 10^{-4}$$

## Question 12.2:

Molar volume is the volume occupied by 1 mol of any (ideal) gas at standard temperature and pressure (STP: 1 atmospheric pressure,  $0^\circ \text{C}$ ). Show that it is 22.4 litres.

## Answer 12.2:

The ideal gas equation relating pressure ( $P$ ), volume ( $V$ ), and absolute temperature ( $T$ ) is given as:

$$PV = nRT$$

Where,

$R$  is the universal gas constant =  $8.314 \text{ J mol}^{-1} \text{ K}^{-1}$

$n$  = Number of moles = 1

$T$  = Standard temperature =  $273 \text{ K}$

$P$  = Standard pressure =  $1 \text{ atm} = 1.013 \times 10^5 \text{ Nm}^{-2}$

$$\therefore V = \frac{nRT}{P} = \frac{1 \times 8.314 \times 273}{1.013 \times 10^5}$$

$$= 0.0224 \text{ m}^3 = 22.4 \text{ litres}$$

Hence, the molar volume of a gas at STP is 22.4 litres.

## Question 12.3:

Figure 12.8 shows plot of  $PV/T$  versus  $P$  for  $1.00 \times 10^{-3} \text{ kg}$  of oxygen gas at two different temperatures.

(a) What does the dotted plot signify?

(b) Which is true:  $T_1 > T_2$  or  $T_1 < T_2$ ?

(c) What is the value of  $PV/T$  where the curves meet on the y-axis?

(d) If we obtained similar plots for  $1.00 \times 10^{-3} \text{ kg}$  of hydrogen, would we get the same value of  $PV/T$  at the point where the curves meet on the y-axis? If not, what mass of hydrogen yields the same value of  $PV/T$  (for low pressure high temperature region of the plot)?

(Molecular mass of  $\text{H}_2 = 2.02 \text{ u}$ , of  $\text{O}_2 = 32.0 \text{ u}$ ,  $R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$ .)



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