Physics

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(Chapter 12) (Kinetic Theory) (Exercises)
(Class 11)

Question 12.1:

Estimate the fraction of molecular volume to the actual volume occupied by oxygen gas at STP. Take the diameter of an oxygen molecule to be 3 Å.

Answer 12.1:

Diameter of an oxygen molecule, $d = 3\text{\AA}$

Radius,
$$r = \frac{d}{2} = \frac{3}{2}$$
 1.5 Å = 1.5 × 10⁻⁸ cm

Actual volume occupied by 1 mole of oxygen gas at STP = 22400 cm³

Molecular volume of oxygen gas,
$$V = \frac{4}{3}\pi r^3$$
.N

Where, N is Avogadro's number = 6.023×10^{23} molecules/mole

$$\therefore V = \frac{4}{3} \times 3.14 \times (1.5 \times 10^{-0})^3 \times 6.023 \times 10^{23} = 8.51 \text{ cm}^3$$

Ratio of the molecular volume to the actual volume of oxygen = $\frac{8.51}{22400}$ = 3.8 × 10⁻⁴

Question 12.2:

Molar volume is the volume occupied by 1 mol of any (ideal) gas at standard temperature and pressure (STP: 1 atmospheric pressure, 0 °C). Show that it is 22.4 litres.

__Answer 12.2:

The ideal gas equation relating pressure (P), volume (V), and absolute temperature (T) is given as: PV = nRT

Where,

R is the universal gas constant = 8.314 J mol-1 K-1

$$n = \text{Number of moles} = 1$$

$$P = \text{Standard pressure} = 1 \text{ atm} = 1.013 \times 105 \text{ Nm}^{-2}$$

$$\therefore V = \frac{nRT}{P} = \frac{1 \times 8.314 \times 273}{1.013 \times 10^5}$$

$$= 0.0224 \, m^3 = 22.4 \, \text{litres}$$

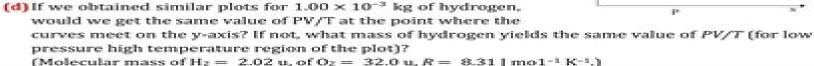
Hence, the molar volume of a gas at STP is 22.4 litres.

Ouestion 12.3

Figure 12.8 shows plot of PV/T versus P for 1.00×10^{-3} kg of oxygen gas at two different temperatures.



(b) Which is true:
$$T_1 > T_2$$
 or $T_1 < T_2$?



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