8 Structure and nomenclature of organic compounds

Unit	4	How are organic compounds categorised, analysed and used?		
Area of Study 1		How can the diversity of carbon compounds be explained and categorised?		
Topic	8	Structure and nomenclature of organic compounds		
Subtopic	8.2	The carbon atom		

Summary

- Carbon atoms have 4 valence electrons and make 4 covalent bonds with other carbon atoms or atoms of different elements.
- Bond energy is the amount of energy required to break the bonds of reactants or the amount of energy released per mole when a covalent bond forms between 2 molecules.
 It is typically measure in kj mol⁻¹.
- The bond length is the distance between the nuclei of the atoms sharing the electrons.
- Carbon molecules with single bonds typically form a tetrahedral arrangement.
- Double C=C bonds are stronger than C−C single bonds and forms flat molecules. C=C bonds are stronger again and form linear molecules.



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