

CHAPTER 7

STUDY GUIDE

Ionic Compounds and Metals

Section 7.1 Ion Formation

In your textbook, read about chemical bonds and formation of ions.

Use each of the terms below just once to complete the passage.

chemical bond	electrons	energy level	ions	noble gases
nucleus	octet	pseudo-noble gas formations	valence	

The force that holds two atoms together is called a(n) (1) _____.
Such an attachment may form by the attraction of the positively charged
(2) _____ of one atom for the negatively charged
(3) _____ of another atom, or by the attraction of charged atoms, which are called
(4) _____. The attractions may also involve
(5) _____ electrons, which are the electrons in the outermost
(6) _____. The (7) _____ are a family of elements that have very
little tendency to react. Most of these elements have a set of eight outermost electrons, which is called a
stable (8) _____. The relatively stable electron structures developed by loss of
electrons in certain elements of groups 3, 4, 13, and 14 are called (9) _____.

For each statement below, write *true* or *false*.

- _____ 10. A positively charged ion is called an anion.
- _____ 11. Elements in group 1 lose their one valence electron, forming an ion with a 1+ charge.
- _____ 12. Elements tend to react so that they acquire the electron structure of a halogen.
- _____ 13. A sodium atom tends to lose one electron when it reacts.
- _____ 14. The electron structure of a zinc ion (Zn^{2+}) is an example of a pseudo-noble gas formation.
- _____ 15. A Cl^- ion is an example of a cation.
- _____ 16. The ending *-ide* is used to designate an anion.
- _____ 17. Nonmetals form a stable outer electron configuration by losing electrons and becoming anions.

Study Guide For Content Mastery Ionic Compounds

Daniela Niemeyer



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Table of Contents Study Guide For Content Mastery Ionic Compounds

1. Understanding the eBook Study Guide For Content Mastery Ionic Compounds
 - The Rise of Digital Reading Study Guide For Content Mastery Ionic Compounds
 - Advantages of eBooks Over Traditional Books
2. Identifying Study Guide For Content Mastery Ionic Compounds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Study Guide For Content Mastery Ionic Compounds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Study Guide For Content Mastery Ionic Compounds
 - Personalized Recommendations
 - Study Guide For Content Mastery Ionic Compounds User Reviews and Ratings
 - Study Guide For Content Mastery Ionic Compounds and Bestseller Lists

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 - Highlighting and Note-Taking Study Guide For Content Mastery Ionic Compounds
 - Interactive Elements Study Guide For Content Mastery Ionic Compounds
8. Staying Engaged with Study Guide For Content Mastery Ionic Compounds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Study Guide For Content Mastery Ionic Compounds
9. Balancing eBooks and Physical Books Study Guide For Content Mastery Ionic Compounds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Study Guide For Content Mastery Ionic Compounds
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Study Guide For Content Mastery Ionic Compounds
 - Setting Reading Goals Study Guide For Content Mastery Ionic Compounds
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Study Guide For Content Mastery Ionic Compounds
 - Fact-Checking eBook Content of Study Guide For Content Mastery Ionic Compounds
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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